

COURSE SCHEME
FOR
B.Sc.(Non-Medical)
2023

FOR BATCH 2023 AND ONWARDS
SARDAR BEANT SINGH STATE UNIVERSITY GURDASPUR

Department of Applied Physics
B.Sc. Non-Medical

1st Semester

Contact hours: 34

Course Code	Course Name	Load Allocated			Credits	Type of Course
		L	T	P		
BSNM-23101	Organic Chemistry	4	0	0	4	Theory
BSNM-23102	Inorganic Chemistry	3	0	0	3	Theory
BSNM-23103	Mathematical Physics	4	0	0	4	Theory
BSNM-23104	Mechanics-I	4	0	0	4	Theory
BSNM-23105	Differential Calculus	4	0	0	4	Theory
BSNM-23106	Solid Geometry	3	0	0	3	Theory
BSNM-23107	English	2	0	0	2	Theory
BSNM-23108	Punjabi-I	2	0	0	2	Theory
BSNM-23109	Chemistry Lab-I	0	0	4	2	Practical
BSNM-23110	Physics Lab-I	0	0	4	2	Practical
	Total	26	0	8	30	

Marks Distribution:

Theory Exam: Internal-40 Marks

External-60 Marks

Practical Exam: Internal-60 Marks

External-40 Marks

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SARDAR BEANT SINGH STATE UNIVERSITY GURDASPUR

Department of Applied Physics
B.Sc. Non-Medical

2nd Semester

Contact hours: 35

Course Code	Course Name	Load Allocated			Credits	Type of Course
		L	T	P		
BSNM-23201	Inorganic Chemistry-II	3	0	0	3	Theory
BSNM-23202	Physical Chemistry-I	3	0	0	3	Theory
BSNM-23203	Mechanics-II	4	0	0	4	Theory
BSNM-23204	Electricity and Magnetism	4	0	0	4	Theory
BSNM-23205	Integral Calculus	4	0	0	4	Theory
BSNM-23206	Theory of equations	3	0	0	3	Theory
BSNM-23207	English-II	2	0	0	2	Theory
BSNM-23208	Punjabi-II	2	0	0	2	Theory
BSNM-23209	Chemistry Lab-II	0	0	4	2	Practical
BSNM-23210	Physics Lab-II	0	0	4	2	Practical
BSNM-23211	Computer Algebra system:MATLAB	0	0	2	1	Practical
	Total	25	0	10	30	

Marks Distribution:

Theory Exam: Internal-40 Marks

External-60 Marks

Practical Exam: Internal-60 Marks

External-40 Marks

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Department of Applied Physics
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3rd Semester

Contact hours: 33

Course Code	Course Name	Load Allocated			Credits	Type of Course
		L	T	P		
BSNM-23301	Organic Chemistry-II	3	0	0	3	Theory
BSNM-23302	Physical Chemistry-II	3	0	0	3	Theory
BSNM-23303	Optics	3	0	0	3	Theory
BSNM-23304	Thermal Physics	3	0	0	3	Theory
BSNM-23305	Analysis-I	3	0	0	3	Theory
BSNM-23306	Differential Equations	3	0	0	3	Theory
BSNM-23307	Statics and dynamics	3	0	0	3	Theory
BSNM-23308	English-III	2	0	0	2	Theory
BSNM-23309	Punjabi-III	2	0	0	2	Theory
BSNM-23310	Chemistry Lab-III	0	0	4	2	Practical
BSNM-23311	Physics Lab-III	0	0	4	2	Practical
	Total	25	0	8	29	

Marks Distribution:

Theory Exam: Internal-40 Marks

External-60 Marks

Practical Exam: Internal-60 Marks

External-40 Marks

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4th Semester

Contact hours: 34

Course Code	Course Name	Load Allocated			Credits	Type of Course
		L	T	P		
BSNM-23401	Inorganic Chemistry-III	3	0	0	3	Theory
BSNM-23402	Organic Chemistry-III	3	0	0	3	Theory
BSNM-23403	Wave Vibrations	4	0	0	4	Theory
BSNM-23404	Electronics	3	0	0	3	Theory
BSNM-23405	Analysis-II	3	0	0	3	Theory
BSNM-23406	Linear Algebra	4	0	0	4	Theory
BSNM-23407	English-IV	2	0	0	2	Theory
BSNM-23408	Punjabi-IV	2	0	0	2	Theory
BSNM-23409	Chemistry Lab-IV	0	0	4	2	Practical
BSNM-23410	Physics Lab-IV	0	0	4	2	Practical
BSNM-23411	Mathematica Software	0	0	2	1	Practical
	Total	24	0	10	29	

Marks Distribution:

Theory Exam: Internal-40 Marks External-60 Marks
Practical Exam: Internal-60 Marks External-40 Marks

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Department of Applied Physics

B.Sc. Non-Medical

5th Semester

Contact hours: 35

Course Code	Course Name	Load Allocated			Credits	Type of Course
		L	T	P		
BSNM-23501	Inorganic Chemistry-IV	3	0	0	3	Theory
BSNM-23502	Physical Chemistry-III	4	0	0	4	Theory
BSNM-23503	Elements of Modern Physics	4	0	0	4	Theory
BSNM-23504	Quantum Mechanics	4	0	0	4	Theory
BSNM-23505	Theory of probability	3	0	0	3	Theory
BSNM-23506	Numerical Analysis	3	0	0	3	Theory
BSNM-23507	English-V	2	0	0	2	Theory
BSNM-23508	Punjabi-V	2	0	0	2	Theory
BSNM-23509	Drug Abuse-I (Problem and Management)	2	0	0	-	Theory
BSNM-23510	Chemistry Lab-V	0	0	4	2	Practical
BSNM-23511	Physics Lab-V	0	0	4	2	Practical
	Total	27	0	8	29	

Marks Distribution:

Theory Exam: Internal-40 Marks

External-60 Marks

Practical Exam: Internal-60 Marks

External-40 Marks

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Department of Applied Physics

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6th Semester

Contact hours: 35

Course Code	Course Name	Load Allocated			Credits	Type of Course
		L	T	P		
BSNM-23601	Organic Chemistry-IV	4	0	0	4	Theory
BSNM-23602	Physical Chemistry-IV	3	0	0	3	Theory
BSNM-23603	Solid State Physics	4	0	0	4	Theory
BSNM-23604	Nuclear and Particle Physics	4	0	0	4	Theory
BSNM-23605	Modern algebra	4	0	0	4	Theory
BSNM-23606	English-VI	2	0	0	2	Theory
BSNM-23607	Environment Science	2	0	0	2	Theory
BSNM-23608	Punjabi- VI	2	0	0	2	Theory
BSNM-23609	Drug Abuse-II (Management and Prevention)	2	0	0	-	Theory
BSNM-23610	Chemistry Lab- VI	0	0	4	2	Practical
BSNM-23611	Physics Lab- VI	0	0	4	2	Practical
	Total	27	0	8	29	

Marks Distribution:

Theory Exam: Internal-40 Marks

External-60 Marks

Practical Exam: Internal-60 Marks

External-40 Marks

BSNM-23101
ORGANIC CHEMISTRY

L T P C
4 0 0 4

Course Objective

To teach the basic principles, reaction mechanisms and stereochemistry of organic compounds. To impart knowledge regarding physical properties and chemical reactions of alkanes, alkenes, dienes, alkynes, arenes, alkyl and aryl halides etc.

I. Structure and Bonding (10)

Hybridization, bond lengths, bond angles, bond energy, localized and delocalized chemical bond, van der Waals interactions, inclusion compounds, clathrates, charge transfer complexes resonance, hyperconjugation, aromaticity, inductive and field effects, hydrogen bonding.

Mechanism of Organic Reactions

Curved arrow notation, drawing electron movements with arrows, half-headed and double-headed arrows, homolytic and heterolytic bond breaking. Types of reagents- electrophiles and nucleophiles. Types of organic reactions. Energy considerations. Reactive intermediates (carbocations, carbanions, free radicals, carbenes, arynes and nitrenes). Assigning formal charges on intermediates and other ionic species.

Methods of determination of reaction mechanism (product analysis, intermediates, isotope effects, kinetic and stereochemical studies).

II. Stereochemistry of Organic Compounds (10)

Isomerism and its types, Optical isomerism - elements of symmetry, molecular chirality, enantiomers, stereogenic center, optical activity, properties of enantiomers, chiral and achiral molecules with two stereogenic centers, diastereomers, threo and erythro, diastereomers, meso compounds, resolution of enantiomers, inversion, retention and racemization. Relative and absolute configuration, sequence rules, D & L and R & S systems of nomenclature.

Geometric isomerism - determination of configuration of geometric isomers. E & Z system of nomenclature, geometric isomerism in oximes and alicyclic compounds. Conformational isomerism - conformational analysis of ethane and n-butane; conformational analysis of cyclohexane, axial and equatorial bonds, conformation of mono substituted cyclohexane derivative.

Newman projection and Sawhorse formulae, Fischer and flying wedge formulae.

Difference between configuration and conformation.

III. Alkanes and Cycloalkanes (10)

Introduction, IUPAC nomenclature, Isomerism and classification of carbon atoms of alkanes. Sources, methods of formation (with special reference to Wurtz reaction, Kolbe reaction, Corey-House reaction and decarboxylation of carboxylic acids). Physical properties and chemical reactions of alkanes.

Mechanism of free radical halogenation of alkanes: orientation, reactivity and selectivity. Cycloalkanes - nomenclature, methods of formation, chemical reactions, Baeyer's strain theory

and its limitations. Ring strain in small rings (cyclopropane and cyclobutane), theory of strainless rings. The case of cyclopropane ring; banana bonds.

IV. Alkenes, Cycloalkenes, Dienes and Alkynes (10)

Alkenes Nomenclature, methods of synthesis (mechanisms of dehydration of alcohols and dehydrohalogenation of alkyl halides, regioselectivity in alcohol dehydration. Saytzeff rule, Hofmann elimination), physical properties and relative stabilities of alkenes. Chemical reactions of alkenes - mechanisms involved in hydrogenation, electrophilic and free radical additions, Markownikoff's rule, hydroboration-oxidation, oxymercuration-reduction. Epoxidation, ozonolysis, hydration, hydroxylation and oxidation with KMnO_4 , Polymerization of alkenes. Substitution at the allylic and vinylic positions of alkenes. Industrial applications of ethylene and propene.

Cycloalkenes Methods of formation, conformation and Chemical reactions of cycloalkenes. *Dienes* Nomenclature and classification of dienes: isolated, conjugated and cumulated dienes. Structure of allenes and butadiene, methods of formation, polymerization. Chemical reactions – 1, 2 and 1,4 addition, Diels-Alder reaction.

Alkynes Nomenclature, structure and bonding in alkynes. Methods of formation. Chemical reactions of alkynes, acidity of alkynes. Mechanism of electrophilic and nucleophilic addition reactions, hydroboration oxidation, metal-ammonia reductions, oxidation and polymerization.

Course Outcomes

1. This course will equip the students with the necessary chemical knowledge concerning the fundamentals in the basic areas of organic chemistry.
2. The students will be able know about the various reactions involved in alkanes, alkenes, alkynes, cycloalkanes, dienes and cycloalkenes.
3. To develop skills to evaluate, analyze and solve problems competently.

Text Books

1. Organic Chemistry, Morrison and Boyd, Prentice- Hall.
2. Fundamentals of Organic Chemistry, Solomons, John Wiley.
3. Organic Chemistry. F.A. Carey, McGraw Hill, Inc.
4. Organic Chemistry, L.G. Wade Jr. Prentice Hall.
5. Organic Chemistry Vol. I, II & III, S.M. Mukherji, S.P. Singh and R.P. Kapoor, Wiley Eastern Ltd (New Age International).
6. Introduction to organic chemistry, Stritwieser, Heathcock and Kosover, Macmillan

BSNM-23102
INORGANIC CHEMISTRY

L T P C
3 0 0 3

Course objective

The objective of this course is to explain the structure and bonding in molecules / ions and predict the structure of molecules / ions. To understand the principles of VSEPR and MOT theory completely. To study the periodical properties of elements and know the structures of ionic solids.

I. Atomic Structure (9)

de Broglie equation, Heisenberg's Uncertainty Principle and its significance. Schrödinger's wave equation and its derivation, significance of ψ and ψ^2 . Quantum numbers. Normalized and orthogonal wave functions. Sign of wave functions. Radial and angular wave functions and distribution curves. Shapes of s, p, d and f orbitals. Pauli's Exclusion Principle, Hund's rule of maximum multiplicity, Aufbau principle and its limitations.

II. Chemical Periodicity (8)

Effective nuclear charge, shielding or screening effect (Slater rules), variation of effective nuclear charge in periodic table. Atomic and ionic radii, Ionization enthalpy, Electron gain enthalpy and their trend in groups and periods. Electronegativity and various scales. Variation of electronegativity with bond order, partial charge, hybridization, group electronegativity.

III. Chemical Bonding I (9)

Ionic bond: General characteristics of ionic compounds, size effects, radius ratio rule and its limitations. Efficiency of packing, Hexagonal close packing, Cubic close packing. Structures of different crystal lattices, Sodium chloride, Cesium chloride, Wurtzite, Zinc blende, Fluorite, Rutile, Cristobalite, Nickel arsenide, Pervoskite, Rhenium oxide, Calcium carbide, The calcite and aragonite structures. Born-Landé equation with derivation and importance of Kapustinskii expression for lattice energy. Madelung constant, Born-Haber cycle and its application, Solvation energy.

IV. Chemical Bonding II (9)

Covalent bond: Lewis structure, Valence Bond theory, VSEPR theory (Prediction of structures and variation of bond angles on the basis of VSEPR theory, Shortcomings of VSEPR theory), Hybridization, Molecular orbital theory (LCAO method). Molecular orbital diagrams of diatomic and simple polyatomic molecules (Be_2 , N_2 , O_2 , F_2 , LiH , NO , CO , HCl , NO_2 , BeH_2 , NO_2^-), Formal charge, Covalent character in ionic compounds, polarizing power and polarizability. Fajan's rules and consequences of polarization. Ionic character in covalent compounds (Bond moment, dipole moment, Percentage ionic character) *Metallic Bond:* Valence bond and band theories. Semiconductors and insulators, defects in solids.

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Weak Interactions: van der Waals forces, ion-dipole forces, dipole-dipole interactions, induced dipole interaction, Hydrogen bonding.

Course Outcomes

1. Students should be able to explain different types of bonds formed by the atoms.
2. They should be able to draw the shapes and geometries of various molecules by MOT and VBT.
3. To develop skills to evaluate the structures of ionic solids and defects in solids.
4. The students should be able to know the various forces exist in molecules.

Text Books

1. D.F.C. Shriver, P.W. Atkins and C.H. Langford, Inorganic Chemistry, ELBS Oxford, 1991.
2. J.E. Huheey, E.A. Keiter, R.L. Keiter, Inorganic Chemistry, 4th Ed, Pearson Education, Singapore, 1999.
3. J.D. Lee, Concise Inorganic Chemistry, ELBS, Oxford 1994.

BSNM- 23103
MATHEMATICAL PHYSICS

L	T	P	C
4	0	0	4

Course Objective: The purpose of the course is to introduce students to methods of mathematical physics and to develop required mathematical skills to solve problems in other fields of theoretical physics taught in different courses this class.

I. First Order and Second Order Ordinary Differential equations (10)

First Order Differential Equations and Integrating Factor. Homogeneous Equations with constant coefficients. Wronskian and general solution. Statement of existence and Uniqueness Theorem for Initial Value Problem.

Calculus of functions of more than one variable: Partial derivatives, exact and inexact differentials. Integrating factor, with simple illustration. Constrained Maximization using Lagrange Multipliers.

II. Vector Calculus (10)

Recapitulation of vectors: Properties of vectors under rotations. Scalar product and its invariance under rotations. Vector product, Scalar triple product and their interpretation in terms of area and volume respectively. Scalar and Vector fields.

III. Vector Differentiation (10)

Directional derivatives and normal derivative. Gradient of a scalar field and its geometrical interpretation. Divergence and curl of a vector field. Del and Laplacian operators. Vector identities.

IV Vector Integration (10)

Ordinary Integrals of Vectors. Multiple integrals, Jacobian. Notion of infinitesimal line, surface and volume elements. Line, surface and volume integrals of Vector fields. Flux of a vector field. Gauss' divergence theorem, Green's and Stokes Theorems and their applications.

Course Outcomes

1. To make them solve IVP, constrained maximization problems using the concepts of ODE and PDE.
2. To interpret the role of vector in applications involves area and volume using vector product and scalar triple product respectively.
3. To understand about geometrical interpretation of gradient, divergence, curl, del and Laplacian operator
4. Ordinary & Multiple integrals of vector field and their applications based on divergence,

green's and stokes theorems.

Text Books

1. Mathematical Methods for Physicists, G.B. Arfken, H.J. Weber, F.E. Harris, 2013, 7thEdn., Elsevier.
2. An introduction to ordinary differential equations, E.A. Coddington, 2009, PHI learning.
3. Differential Equations, George F. Simmons, 2007, McGraw Hill.
4. Mathematical Tools for Physics, James Nearing, 2010, Dover Publications.
5. Mathematical methods for Scientists and Engineers, D.A. McQuarrie, 2003, Viva Book.
6. Advanced Engineering Mathematics, D.G. Zill and W.S. Wright, 5 Ed., 2012, Jones and Bartlett Learning.
7. Mathematical Physics, Goswami, 1st edition, Cengage Learning.
8. Engineering Mathematics, S.Pal and S.C. Bhunia, 2015, Oxford University Press.
9. Advanced Engineering Mathematics, Erwin Kreyszig, 2008, Wiley India.

BSNM- 21104 Mechanics-I

L	T	P	C
4	0	0	4

Course Objectives: The objective of this course is to illustrate the laws of motion, kinematics of motion and their interrelationship. Mechanics course help the student to develop this ability to visualize, which is so vital to problem formulation.

I. Fundamentals of Dynamics (15)

Co-ordinate system and Motion of a Particle: Reference frames, Inertial frames; area, volume, displacement, velocity and acceleration in Cartesian and Spherical Polar co-ordinate systems; Introduction to cylindrical coordinate system, Solid angle, Galilean transformations; Galilean invariance of space & time intervals, Review of Newton's Laws of Motion. Dynamics of a system of particles. Centre of Mass. Principle of conservation of momentum. Impulse.

II. Work and Energy (10)

Work and Kinetic Energy Theorem. Conservative and non-conservative forces. Potential Energy. Energy diagram. Stable and unstable equilibrium. Elastic potential energy. Force as gradient of potential energy. Work & Potential energy. Work done by non-conservative forces. Law of conservation of Energy.

Elastic and Inelastic Scattering

Types of scattering and conservation laws, Laboratory and centre of mass systems, collision of particles which stick together, General elastic collision of particles of different mass, Cross-section of elastic scattering, Rutherford scattering.

III. Rotational Dynamics (10)

Angular momentum of a particle and system of particles. Torque. Principle of conservation of angular momentum. Rotation about a fixed axis. Moment of Inertia. Calculation of moment of inertia for rectangular, cylindrical and spherical bodies. Kinetic energy of rotation. Motion involving both translation and rotation. Cylinder on an accelerated rough plane, Behavior of angular momentum vector, Principal axes and Euler's equations, Elementary Gyroscope, Motion of a spinning top.

IV. Elasticity (10)

Hooke's law-Stress-strain diagram-Elastic moduli-Relation between elastic constants- Poisson's Ratio-Expression for Poisson's ratio in terms of elastic constants - Work done in stretching and work done in twisting a wire-Twisting couple on a cylinder - Determination of Rigidity modulus by static torsion - Torsional pendulum-Determination of Rigidity modulus and moment of inertia - q , η , and σ by Searles method.

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Suggested Books:

1. An introduction to mechanics, D. Kleppner, R.J. Kolenkow, 1973, McGraw-Hill.
2. Mechanics, Berkeley Physics, vol.1, C.Kittel, W.Knight, et.al. 2007, Tata McGraw-Hill.
3. Physics, Resnick, Halliday and Walker 8/e. 2008, Wiley.
4. Analytical Mechanics, G.R. Fowles and G.L. Cassiday. 2005, Cengage Learning.
5. Feynman Lectures, Vol. I, R.P .Feynman, R.B.Leighton, M.Sands, 2008, PearsonEducation
6. Introduction to Special Relativity, R. Resnick, 2005, John Wiley and Sons.
7. University Physics, Ronald Lane Reese, 2003, Thomson Brooks/Cole.
8. Mechanics, D.S. Mathur, S. Chand and Company Limited, 2000
9. University Physics. F.W Sears, M.W Zemansky, H. D Young 13/e, 1986, AddisonWesley.
10. Physics for scientists and Engineers with Modern Phys., J.W. Jewett, R.A. Serwa, 2010,Cengage Learning.
11. Theoretical Mechanics, M.R. Spiegel, 2006, Tata McGraw Hill.

BSNM-23105
DIFFERENTIAL CALCULUS

L	T	P	C
4	0	0	4

Course Objective:

The main objective of the course is to demonstrate an understanding of the foundations and history of differential calculus.

I. Definition of a sequence. (10)

Limit of a sequence, theorems on limits of sequences, bounded, and monotonic sequences. Least upper bound and greatest lower bound of a sequence. Limit superior, limit inferior. Nested Intervals. Cauchy's convergence criterion, infinite series.

II. Limits of Functions (10)

ϵ - δ definition, right- and left-hand limits. Theorems on limits. Infinity. Special Limits. Continuity, ϵ - δ definition, right- and left-hand Continuity, continuity in an interval, theorems on continuity, piecewise continuity, uniform Continuity.

III. Derivatives (10)

The concept and definition of a derivative, right- and left-hand derivatives, differentiability in an interval, piecewise differentiability, differentials, differentiation of composite functions, implicit differentiation, mean value theorems, Taylor theorem, applications.

IV. (10)

Functions of two or more variables, neighborhoods, regions, limits, iterated limits, continuity, uniform continuity, partial derivatives, higher-order partial derivatives, differentials, theorems on differentials, differentiation of composite functions, Euler's theorem on homogeneous functions. Implicit functions, Jacobians, partial derivatives using Jacobians, theorems on Jacobians, applications.

Course Outcomes: After completion of the course, students will be able to

1. Verify the value of the limit of a function at a point using the definition of the limit,
2. Know the difference between sequence and series,
3. Find the extreme values of function

Reference Books

1. Robert Wrede & Murray R. Spiegel, Advanced Calculus, 3rd Ed, Schaum's Outline Series (McGraw Hill), 2010.

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2. Maurice D Weir, Frank R. Giordano and Joel Hass, Thomas' Calculus, 11th Ed, Pearson, 2008.
3. James Stewart, Calculus, 5th Edition, Brooks/Cole(Thomson), 2003.
4. Theoretical Mechanics, M.R. Spiegel, 2006, Tata McGraw Hill.

BSNM-23106
SOLID GEOMETRY

L	T	P	C
3	0	0	3

Course Objective

To enable the students to visualize shapes of solid geometrical objects well, to develop visual spatial skills in solid geometry learning and to enable the students to express the solid shapes in mathematical expressions.

I. (9)

The concept of co-ordinates, co-ordinate of a point in space, distance between two points. Plane: Definition of a plane, Normal form of the equation of a plane, Transformation from general form to normal form, Equation of plane in terms of its intercepts on the axis, Equations of the plane through the given points, Length of the perpendicular from a given point to a given plane, Bisectors of angles between two planes, Combined equation of two planes, Orthogonal projection on a plane.

II. Sphere (9)

Definition and equation of the sphere; Equation of the sphere through four given points; Plane sections of a sphere; Intersection of two spheres; Equation of a circle; Sphere through a given circle; Intersection of a sphere and a line; Power of a point; Tangent plane; Plane of contact; Polar plane; Pole of a plane; Conjugate points; Conjugate planes; Angle of intersection of two spheres; Conditions for two spheres to be orthogonal; Radical plane; Coaxial system of spheres.

III. Cone (9)

Definitions of a cone; vertex; guiding curve; generators; Equation of the cone with a given vertex and guiding curve; Enveloping cone of a sphere; Equations of cones with vertex at origin are homogenous; Condition that the general equation of the second degree should represent a cone; Condition that a cone may have three mutually perpendicular generators; Intersection of a line and a quadric cone; Tangent lines and tangent plane at a point; Condition that a plane may touch a cone; Reciprocal cones; Intersection of two cones with a common vertex; Right circular cone; Equation of the right circular cone with a given vertex; axis and semi-vertical angle.

IV. Cylinder (8)

Definition of a cylinder, Equation to the cylinder whose generators intersect a given conic and are parallel to a given line; Enveloping cylinder of a sphere; The right circular cylinder; Equation of the right circular cylinder with a given axis and radius.

Course Outcomes:

Upon successful completion of this course, students will:

1. be able to visualize shapes of solid geometry objects well.

2. be able to express the solid shapes in mathematical expressions.
3. develop visual spatial skills in solid geometry learning.

Text Books

1. Shanti Narayan and P. K. Mittal, Analytical Solid Geometry, 17th Edition, S. Chand & Company, 2007.
2. P. K. Jain, A Textbook of Analytical Geometry of Three Dimensions, New Age International, 2005.

BSNM-23107

English

L	T	P	C
2	0	0	2

Course Objective

To help the students become proficient in LSRW-Listening, Speaking, Reading & Writing skills. To help the students become the independent users of English language. It educates their emotions and enhances their power of imagination.

I. Literature

(6)

The Poetic Palette (Orient BlackSwan, Second Edition, 2016)

The following poems from this anthology are prescribed:

1. Apparently With No Surprise: Emily Dickinson
2. Fool and Flea: Jeet Thayil
3. The Soul's Prayer: Sarojini Naidu
4. I Sit and Look Out: Walt Whitman
5. Women's Rights: Annie Louise Walker
6. Pippa's Song: Robert Browning

Vocabulary

Antonyms; Synonyms; One-word substitution; Homophones/Homonyms; Abbreviations

II. Literature

(7)

(b) *Prose Parables (Orient Black Swan, 2013)*

The following stories from the above volume are prescribed:

- a. The Eyes Are Not Here: Ruskin Bond
- b. Grief: Anton Chekov
- c. The Doctor's Word: R.K. Narayan
- d. The Doll's House: Katherine Mansfield
- e. Dusk: H.H. Munroe (Saki)
- f. The Kabuli wallah : Rabindranath Tagore

Grammar

Parts of Speech; Articles, Determiners; Modals; Modifiers; Prepositions; Voice; Transformation of sentences.

III. Close Reading

(6)

Comprehension; Summarizing; Paraphrasing; Analysis and Interpretation; Translation (from Hindi/Punjabi to English and vice-versa)

IV.

(6)

Essay Writing -Descriptive/Narrative/Argumentative; Business letters; Précis Writing

Course Outcomes

1. Students will heighten their awareness of correct usage of English grammar in writing and speaking.
2. They will enlarge their vocabulary by keeping a vocabulary journal
3. They will strengthen their ability to write academic papers, essays and summaries using the process approach.

Text Books

1. Oxford Practice Grammar by John Eastwood (Ed. 2014)
2. Business English, Pearson, 2008.
3. Language, Literature and Creativity, Orient Black swan, 2013.
4. Language through Literature (forthcoming) ed. Dr. Gauri Mishra, Dr. Ranjana Kaul, Dr. Brati Biswas
5. Study Writing. Liz Hamp-Lyons and Ben Heasley. Cambridge University Press. 2006.

BSNM-23108
Punjabi-I

L T P C
2 0 0 2

ਵਿਸ਼ਾ ਉਦੇਸ਼: ਇਹ ਵਿਸ਼ਾ ਪੜ੍ਹਾਉਣ ਦਾ ਮੂਲ ਉਦੇਸ਼ ਵਿਦਿਆਰਥੀਆਂ ਨੂੰ ਪੰਜਾਬੀ ਸਾਹਿਤ ਦੇ ਵੱਖ-ਵੱਖ ਰੂਪਾਂ (ਕਵਿਤਾ, ਕਹਾਣੀ, ਲੇਖ, ਪ੍ਰਮੁੱਖ ਸਖਸ਼ੀਅਤਾਂ ਦੀਆਂ ਜੀਵਨੀਆਂ) ਦੇ ਮਹੱਤਵ ਅਤੇ ਪੰਜਾਬੀ ਵਿਆਕਰਣ ਤੋਂ ਜਾਣੂ ਕਰਵਾਉਣਾ ਹੈ।

ਭਾਗ-ਪਹਿਲਾ (ਕਵਿਤਾ)

ਰਉਂ ਰੁੱਖ (ਭਾਈ ਵੀਰ ਸਿੰਘ)
ਸਮਾਂ (ਭਾਈ ਵੀਰ ਸਿੰਘ)
ਇੱਛਾ ਬਲ ਤੇ ਡੂੰਘੀਆਂ ਛਾਵਾਂ (ਭਾਈ ਵੀਰ ਸਿੰਘ)
ਰਾਧਾ ਸੰਦੇਸ਼ (ਧਨੀ ਰਾਮ ਚਾਤ੍ਰਕ)
ਸਿਦਕਾਂ ਵਾਲਿਆਂ ਦੇ ਬੇੜੇ ਪਾਰ ਨੇ (ਧਨੀ ਰਾਮ ਚਾਤ੍ਰਕ)

ਭਾਗ-ਦੂਜਾ (ਕਹਾਣੀ)

ਭੂਆ (ਨਾਨਕ ਸਿੰਘ)
ਬਾਗੀ ਦੀ ਧੀ (ਗੁਰਮੁਖ ਸਿੰਘ ਮੁਸਾਫ਼ਿਰ)
ਪੇਮੀ ਦੇ ਨਿਆਣੇ (ਸੰਤ ਸਿੰਘ ਸੇਖੋਂ)

ਭਾਗ-ਤੀਜਾ (ਲੇਖ)

ਘਰ ਦਾ ਪਿਆਰ (ਪ੍ਰੋ. ਪੂਰਨ ਸਿੰਘ)
ਮਨ ਦੀ ਮੈਜ (ਗਿ. ਲਾਲ ਸਿੰਘ 'ਕਮਲਾ ਅਕਾਲੀ')
ਗੁਰ-ਸੰਗਤ ਬਾਣੀ (ਗਿ. ਹੀਰਾ ਸਿੰਘ ਦਰਦ)
ਕਾਠ ਦੀ ਰੋਟੀ (ਪ੍ਰੋ. ਸਾਹਿਬ ਸਿੰਘ)

ਭਾਗ-ਚੌਥਾ (ਪ੍ਰਮੁੱਖ ਸਖਸ਼ੀਅਤਾਂ)

ਭਗਵਾਨ ਬੁੱਧ (ਪ੍ਰੋ. ਹਰਨਾਮ ਸਿੰਘ ਸ਼ਾਨ)
ਬਾਬਾ ਫ਼ਰੀਦ (ਅਵਤਾਰ ਸਿੰਘ ਆਜ਼ਾਦ)
ਗੁਰੂ ਨਾਨਕ ਦੇਵ ਜੀ (ਪ੍ਰਿੰ. ਤੇਜਾ ਸਿੰਘ)

ਭਾਗ-ਪੰਜਵਾਂ (ਵਿਆਕਰਣ):

ਭਾਸ਼ਾ ਦਾ ਟਕਸਾਲੀ ਰੂਪ, ਭਾਸ਼ਾ ਅਤੇ ਉੱਪ-ਭਾਸ਼ਾ ਵਿੱਚ ਅੰਤਰ, ਪੰਜਾਬੀ ਦੀਆਂ ਉੱਪ-ਭਾਸ਼ਾਵਾਂ, ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਦਾ ਨਿਕਾਸ ਅਤੇ ਵਿਕਾਸ, ਭਾਸ਼ਾ ਅਤੇ ਲਿੱਪੀ, ਗੁਰਮੁਖੀ ਲਿੱਪੀ ਦਾ ਨਿਕਾਸ, ਵਿਕਾਸ ਅਤੇ ਵਿਸ਼ੇਸ਼ਤਾਵਾਂ।

ਵਿਸ਼ਾ ਨਤੀਜਾ: ਵਿਦਿਆਰਥੀ ਵੱਖ-ਵੱਖ ਕਵਿਤਾਵਾਂ, ਕਹਾਣੀਆਂ, ਲੇਖਾਂ, ਪ੍ਰਮੁੱਖ ਸਖਸ਼ੀਅਤਾਂ ਦੀਆਂ ਜੀਵਨੀਆਂ ਦੇ ਸਾਰ ਅਤੇ ਸਿੱਖਿਆ ਰਾਹੀਂ ਪੰਜਾਬ ਦੇ ਇਤਿਹਾਸ ਅਤੇ ਪੰਜਾਬੀ ਸੱਭਿਆਚਾਰ ਵਾਰੇ ਜਾਣਕਾਰੀ ਪ੍ਰਾਪਤ ਕਰਕੇ ਵਧੀਆ ਜੀਵਨ ਜਾਚ ਗ੍ਰਹਿਣ ਕਰਕੇ ਪੰਜਾਬ ਪ੍ਰਾਂਤ ਅਤੇ ਭਾਰਤ ਦੇਸ਼ ਦੀ ਤਰੱਕੀ ਵਿੱਚ ਵਡਮੁੱਲਾ ਯੋਗਦਾਨ ਪਾਉਣ ਦੇ ਕਾਬਲ ਹੋ ਜਾਣਗੇ। ਪੰਜਾਬੀ ਵਿਆਕਰਣ ਦੀ ਜਾਣਕਾਰੀ ਵਿਦਿਆਰਥੀਆਂ ਨੂੰ ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਪੜ੍ਹਨ, ਲਿਖਣ ਅਤੇ ਸਮਝਣ ਵਿੱਚ ਨਿਪੁੰਨਤਾ ਪ੍ਰਦਾਨ ਕਰੇਗੀ।

ਪੁਸਤਕਾਂ ਦੀ ਸੂਚੀ

1. ਦੇ ਰੰਗ (ਸੰਪਾਦਕ-ਹਰਜਿੰਦਰ ਸਿੰਘ ਢਿੱਲੋਂ ਅਤੇ ਪ੍ਰੀਤਮ ਸਿੰਘ ਸਰਗੋਧਾ, ਪੰਜਾਬੀ ਅਤੇ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ ਵਿਭਾਗ, ਗੁਰੂ ਨਾਨਕ ਦੇਵ ਯੂਨੀਵਰਸਿਟੀ ਅੰਮ੍ਰਿਤਸਰ)।
2. ਆਧੁਨਿਕ ਪੰਜਾਬੀ ਵਾਰਤਕ (ਸੰਪਾਦਕ-ਡਾ. ਗੁਰਬਚਨ ਸਿੰਘ ਤਾਲਿਬ, ਪੰਜਾਬੀ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ ਅੰਮ੍ਰਿਤਸਰ)।
3. ਸੰਸਾਰ ਦੀਆਂ ਪ੍ਰਸਿੱਧ ਹਸਤੀਆਂ (ਪ੍ਰਿੰ. ਤੇਜਾ ਸਿੰਘ ਅਤੇ ਪ੍ਰੋ. ਹਰਨਾਮ ਸਿੰਘ ਸ਼ਾਨ, ਪੰਜਾਬੀ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ ਅੰਮ੍ਰਿਤਸਰ)।

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4. ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਬੋਧ (ਕਸਤੂਰੀ ਲਾਲ ਐਂਡ ਸੰਨਜ਼, ਅੰਮ੍ਰਿਤਸਰ)।
5. ਪੰਜਾਬੀ ਵਿਆਕਰਣ: ਸਿਧਾਂਤ ਤੇ ਵਿਹਾਰ (ਚੇਤਨਾ ਪ੍ਰਕਾਸ਼ਨ, ਲੁਧਿਆਣਾ)।

BSNM-23109
CHEMISTRY LAB- I

L	T	P	C
0	0	4	2

Course Objective: The objective of this course is to provide practical knowledge and illustrative experiments regarding qualitative analysis, isolation, and purification of organic compounds.

Inorganic Chemistry: Semi Micro analysis. Cation analysis, Separation and identification of ions from groups I, II, III, IV, V, and VI. Anionic analysis. Four ions with no interference.

Organic Chemistry Laboratory Techniques:

Determination of Melting Point

Naphthalene 80-82⁰C

Cinnamic acid 132.5-133⁰C

Benzoic acid 121.5-122⁰C

Salicylic acid 157.5-158⁰C

Urea 132.5-133⁰C

Acetanilide 113.5-114⁰C

Succinic Acid 184.5-185⁰C

m-dinitro benzene 90⁰C

p-dichlorobenzene 52⁰C

Aspirin 135⁰C

Determination of Boiling Point

Ethanol 78⁰C

Cyclohexane 81.4⁰C

Benzene 80⁰C

Toluene 110⁰C

BSNM-23110
Physics Lab-I

L	T	P	C
0	0	4	2

Course Objective: The objective of this course is to enable the students to verify some of the concepts learnt in the theory courses. The course provides them training in carrying out precise measurements and handling sensitive equipment.

Students are required to perform at least 06 experiments from the following:

1. Measurements of length (or diameter) using vernier caliper, screw gauge, and travelling microscope. Use of Plumb line and Spirit level.
2. Analysis of experimental data by:
 - (a) fitting the given data to a straight line
 - b) to study probable error in observations.
3. To determine the height of an inaccessible object using a sextant.
4. To determine the horizontal distance of an object using a sextant.
5. To determine the vertical distance of an object using a sextant.
6. To verify the law of vibrating string by Melde's experiment.
7. To setup CRO for Sine and Square wave and to find their frequency and amplitude.
8. To study the Motion of Spring and calculate (a) Spring constant, (b) g and (c) Modulus of rigidity.
9. To establish a relation between angular acceleration α and torque τ , and hence to find out the moment of Inertia of flywheel.
10. Study the dependence of the moment of Inertia on distribution of mass (by noting the time periods of oscillations) using objects of various shape but of same mass.
11. To determine the Young's Modulus of a Wire by Optical Lever Method.
12. To determine the Young's Modulus of a Wire by Searle's method.
13. To determine the Modulus of Rigidity of a Wire by Maxwell's needle.

Text Books

1. Advanced Practical Physics for students, B. L. Flint and H.T. Worsnop, 1971, Asia Publishing House.
2. Advanced level Physics Practicals, Michael Nelson and Jon . Ogborn, 4th Edition, reprinted 1985, Heinemann Educational Publishers
3. Engineering Practical Physics, S.Panigrahi & B.Mallick, 2015, Cengage Learning India Pvt. Ltd.
4. Practical Physics, G.L. Squires, 2015, 4th Edition, Cambridge University Press.
5. B Sc Practical Physics by C. L. Arora, S. Chand & Co.

BSNM-23201
INORGANIC CHEMISTRY-II

L	T	P	C
3	0	0	3

Course Objective

To teach the fundamental concepts of Inorganic Chemistry and their applications. To fully understand the chemistry of transition and inner transition elements. To impart knowledge regarding physical and chemical properties of S, P D block elements.

I. Chemistry of s Block Elements (9)

General characteristics (melting point, flame color, reducing nature, diagonal relationships and anomalous behavior of first member of each group).

Reactions of alkali and alkaline earth metals with oxygen, hydrogen, nitrogen and water. Ease of formation, thermal stability and solubility of the following alkali and alkaline earth metal compounds: hydrides, oxides, peroxides, superoxides, carbonates, nitrates, sulphates.

Complex formation tendency of s-block elements; crown ethers, cryptands and podands of Group I; basic beryllium acetate, beryllium nitrate, EDTA complexes of calcium and magnesium.

Solutions of alkali metals in liquid ammonia and their properties.

II. Chemistry of p Block Elements (12)

Electronic configuration, atomic and ionic size, metallic/non-metallic character, melting point, ionization enthalpy, electron gain enthalpy, electronegativity, inert pair effect, diagonal relationship between B and Si and anomalous behaviour of first member of each group.

Group III (Boron Group): Oxides, halides and hydrides of group III elements, boron sesquioxide and borates structure of borates, trihalides and lower halides of boron, preparation of boron hydrides reactions and structures of boranes.

Group IV (Carbon Group): Structure and allotropy of the elements, types and structure of carbides, oxides of carbon and silicon, types and structures of silicates, Organo – silicon compounds and the silicones, halides of IV group elements.

Group V (Nitrogen Group): Hydrides, properties and structure of ammonia, hydrazine, hydroxylamine, trihalides and Pentahalides of V groups elements, oxides of nitrogen, structure of N₂O, NO, N₂O₃, N₂O₄ and N₂O₅, oxo acids of nitrogen and phosphorous, phosphazenes and cyclophosphazenes.

Group VI (Oxygen Group): Structure and allotropy of the elements. Oxides of sulfur (structure of SO₂ and SO₃) oxoacids of sulfur halides of sulfur, selenium and tellurium, compounds of Sulfur and nitrogen (S₄N₄).

Group VII: Oxides of halogens (OF₂, O₂F₂, Cl₂O, ClO₂, Cl₂O₆, BrO₂, I₂O₅) (structures), Preparation, reaction and structure of interhalogen compounds. (ClF₃, BrF₃, I₂, Cl₅, IF₅, IF₇), Polyhalides, basic properties of halogens.

III. Acids-bases

(7)

Various definitions of acids and bases, A generalized acid-base concept, Measurement of acid-base strength, Lewis interactions in non-polar solvents, Systematics of Lewis acid- base interactions, Bond energies, steric effects, solvation effects and acid-base anomalies, Classification of acids and bases as hard and soft. Pearson's HSAB concept, acid-base strength and hardness and softness. Symbiosis, theoretical basis of hardness and softness, electronegativity and hardness and softness.

IV. Chemistry of Transition Elements

(7)

Characteristic properties of d-block elements. Properties of the elements of the first transition series, their simple compounds and complexes illustrating relative stability of their oxidation states, coordination number and geometry. General characteristics of elements of Second and Third Transition Series, comparative treatment with their 3d analogues in respect of ionic radii, oxidation states, magnetic behavior.

Course Outcome

1. Students should be able to know the anomalous properties and compounds of carbon, Boron & Silicon.
 2. The learners should be able to know the physical and chemical properties of alkali and alkaline earth metals.
 3. They should be able to learn the Different concept of acid and bases.
- Students should be able to learn the complexes of alkali and alkaline earth metals.

Text Books

1. J.D. Lee, Concise Inorganic Chemistry, 4th Ed.
2. J.E. Huheey, Inorganic Chemistry, Harper & Row.
3. F.A.Cotton and G. Wilinon, Advanced Inorganic Chemistry, Interscience Publishers.
4. 4.N.N. Greenwood and A. Earnshaw, Chemistry of Elements, Pergamon Press

BSNM-23202
PHYSICAL CHEMISTRY-I

L	T	P	C
3	0	0	3

Course Objective: This course will equip students with the necessary chemical knowledge concerning the fundamentals in the basic areas of physical chemistry viz. different states of matter. The problem solving skills of students are expected to be enhanced through numerical problems.

I. Gaseous state

(7)

Kinetic molecular theory of gases, derivation of kinetic gas equation, deduction of gas laws from kinetic gas equation, imperfection in real gases, the compressibility of real gases, isotherms of real gases, equations of state, Causes of deviation from ideal

behaviour. van der Waals equation of state, its derivation and application in explaining real gas behaviour, calculation of Boyle temperature. Isotherms of real gases and their comparison with van der Waals isotherms, continuity of states, critical state, relation between critical constants and van der Waals constants, law of corresponding states.

II. Liquids state

(8)

Qualitative treatment of the structure of the liquid state; physical properties of liquids; vapour pressure, surface tension and coefficient of viscosity, and their determination. Effect of addition of various solutes on surface tension and viscosity. Explanation of cleansing action of detergents. Temperature variation of viscosity of liquids and comparison with that of gases.

III. Colloidal State

(10)

Definition of colloids, classification of colloids. Solids in liquids (Sol): kinetic, optical and electrical, properties, stability of colloids, protective action, Hardy Schulze law, gold number. Liquids in liquids (emulsions): Types of emulsions, preparation. Emulsifiers. General applications of colloids.

IV. Solutions, Dilute Solutions and Colligative Properties

(10)

Ideal and non-ideal solutions, methods of expressing concentrations of solutions, activity and activity coefficient. Dilute solution, colligative properties, Raoult's law, relative lowering of vapour pressure, molecular weight determination. Osmosis, Law of osmotic pressure and its measurement, determination of molecular weight from osmotic pressure. Elevation of boiling point and depression of freezing point, Thermodynamic derivation of relation between molecular weight and elevation in boiling point and depression in freezing point. Experimental methods for determining various colligative properties. Abnormal molar mass degree of dissociation and association of solutes.

Course Outcome

By the end of the course, students will be able to:

1. Understand the basic concept of kinetic theory of gases and different aspects of gaseous state.
2. Gain insight into the physical significance of various properties of liquid.
3. Understand various types of colloids and its applications.
4. Know the qualitative properties of solution, the depression in freezing point, elevation in boiling point and osmotic pressure.
5. Derive the thermodynamic relations between the colligative properties and understand their applications in everyday life.

Text Books

1. Principles of physical chemistry, S.H. Maron & C.F. Prutton.
2. Physical Chemistry, K.J. Laidler.
3. Atkins, P. W. & Paula, J. de Atkin's Physical Chemistry Ed., Oxford University Press 13 (2006).
4. Ball, D. W. Physical Chemistry Thomson Press, India (2007).
5. Castellan, G. W. Physical Chemistry 4th Ed. Narosa (2004).
6. Mortimer, R. G. Physical Chemistry 3rd Ed. Elsevier: NOIDA, UP (2009).

BSNM-23203
Mechanics-II

L T P C
4 0 0 4

Course Objective

This course imparts knowledge and help in understanding of the fundamental concepts kinematics of a body. This course also explains the meaning and significance of the postulate of Special Theory of Relativity. Students will be able to solve problems involving the application of the conservation laws of energy and angular momentum. The problem solving skills of students are expected to be enhanced through numerical problems.

I. Gravitation (10)

Law of gravitation. Gravitational potential energy. Inertial and gravitational mass. Potential and field due to spherical shell and solid sphere. Force between a Point Mass and Spherical shell. Force between a Point Mass and Solid Sphere, Gravitational and Electrostatic self-energy. Gravitational energy of the Galaxy and of uniform sphere.

II. Central Force Motion (10)

Motion of a particle under a central force field. Two-body problem and its reduction to one-body problem and its solution. The energy equation and energy diagram. Kepler's Laws. Satellite in circular orbit and applications.

Geosynchronous orbits. Basic idea of global positioning system (GPS).

Non-Inertial Systems: Non-inertial frames and fictitious forces. Uniformly rotating frame. Laws of physics in rotating coordinate systems. Centrifugal force. Coriolis force and its applications. Components of velocity and acceleration in cylindrical and spherical Coordinate systems.

III. Oscillations (10)

Simple Harmonic Oscillations (SHM). Differential equation of SHM and its solution. Kinetic energy, potential energy, total energy and their time-average values. Damped oscillation. Forced oscillations: Transient and steady states; Resonance, sharpness of resonance; power dissipation and Quality Factor.

IV. Special Theory of Relativity (10)

Michelson-Morley Experiment. Postulates of Special Theory of Relativity. Lorentz Transformations. Simultaneity and order of events. Lorentz contraction. Time dilation. Relativistic transformation of velocity, frequency and wave number. Relativistic addition of velocities. Variation of mass with velocity. Massless Particles. Mass-energy Equivalence. Relativistic Kinematics. Transformation of Energy and Momentum.

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Course Outcome: On the successful completion of the course, students will be able to

1. Understand and define the laws involved in mechanics
2. Gain deeper understanding of mechanics and its fundamental concepts
3. Provide the students with an idea of relativity which are essential tools in problem solving.

Text Books

1. An introduction to mechanics, D. Kleppner, R.J. Kolenkow, 1973, McGraw-Hill.
2. Mechanics, Berkeley Physics, vol.1, C.Kittel, W.Knight, et.al. 2007, Tata McGraw-Hill.
3. Physics, Resnick, Halliday and Walker 8/e. 2008, Wiley.
4. Analytical Mechanics, G.R. Fowles and G.L. Cassiday. 2005, Cengage Learning.
5. Feynman Lectures, Vol. I, R.P .Feynman, R.B.Leighton, M.Sands, 2008, PearsonEducation
6. Introduction to Special Relativity, R. Resnick, 2005, John Wiley and Sons.
7. University Physics, Ronald Lane Reese, 2003, Thomson Brooks/Cole.
8. Mechanics, D.S. Mathur, S. Chand and Company Limited, 2000
9. University Physics. F.W Sears, M.W Zemansky, H. D Young 13/e, 1986, Addison Wesley.
10. Physics for scientists and Engineers with Modern Phys., J.W. Jewett, R.A. Serwa, 2010, Cengage Learning
11. Theoretical Mechanics, M.R. Spiegel, 2006, Tata McGraw Hill.

BSNM-23204
ELECTRICITY AND MAGNETISM

L	T	P	C
4	0	0	4

Course Objectives

The aim of this course is to acquire in-depth knowledge in electrostatics and magnetostatics so that students would apply theories of static and moving charges and extend its applications to instruments involving electric and magnetic fields. This course will also give idea on the fundamentals of electromagnetic conduction and electromagnetic waves..

I. Electrostatics and Dielectrics (10)

Electrostatic Field, electric flux, Gauss's theorem of electrostatics. Applications of Gauss theorem- Electric field due to point charge, infinite line of charge, uniformly charged spherical shell and solid sphere, plane charged sheet, charged conductor. Electric potential as line integral of electric field, potential due to a point charge, electric dipole, uniformly charged spherical shell and solid sphere. Calculation of electric field from potential. Capacitance of an isolated spherical conductor. Parallel plate, spherical and cylindrical condenser. Energy per unit volume in electrostatic field. Dielectric medium, Polarisation, Displacement vector. Gauss's theorem in dielectrics. Parallel plate capacitor completely filled with dielectric.

II. Magnetism (10)

Magnetostatics: Biot-Savart's law & its applications- straight conductor, circular coil, solenoid carrying current. Divergence and curl of magnetic field. Magnetic vector potential. Ampere's circuital law.

Magnetic properties of materials: Magnetic intensity, magnetic induction, permeability, magnetic susceptibility. Brief introduction of dia-,para- and ferro-magnetic materials.

III. Electromagnetic Induction (10)

Faraday's laws of electromagnetic induction, Lenz's law, self and mutual inductance, L of single coil, M of two coils. Energy stored in magnetic field.

IV. Maxwell's equations and Electromagnetic wave propagation (10)

Equation of continuity of current, Displacement current, Maxwell's equations, Poynting vector, energy density in electromagnetic field, electromagnetic wave propagation through vacuum and isotropic dielectric medium, transverse nature of EM waves, polarization.

Course Outcome: On the successful completion of the course, students will be able to

1. Recognize basic terms in electricity and magnetism
2. Understand the laws of electrostatics and magnetostatics

Text Books

1. Edward M. Purcell, Electricity and Magnetism, McGraw-Hill Education 1986.

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- 2 J.H. Fewkes & J. Yarwood. Electricity and Magnetism, Oxford Univ. Press Vol. I, 1991.
3. D C Tayal, Electricity and Magnetism, Himalaya Publishing House 1988.
4. Ronald Lane Reese, University Physics, Thomson Brooks/Cole 2003.
5. D.J. Griffiths, Introduction to Electrodynamics, Benjamin Cummings 3rd Edn, 1998

BSNM-23205
INTEGRAL CALCULUS

L	T	P	C
4	0	0	4

Course Objectives:

The objectives of this course are to make the students understand the concept of integration and its definition as limit of sum and area under curve. The concept of improper integrals will also be taught.

I. (10)
Integrals of functions of one variable, geometrical interpretation of integral as area, integration of standard functions, integration by substitution and parts, Integration by Partial fractions, integration of rational and irrational functions. Properties of definite integrals.

II. (10)
Reduction formulae for integrals of rational, trigonometric, exponential and logarithmic functions and of their combinations. Areas and lengths of curves in the plane, volumes and surfaces area of solids of revolution.

III. (10)
Integrals of functions of two variables, double integrals, Applications to evaluation of area, volumes and surfaces of solids of revolution, Change of order of Integration. Change of variables.

IV. (10)
Integrals of functions of three variables, Triple integral, Evaluation of volume, density etc., Change of order of Integration. Change of variables. Implicit and Explicit functions, Integration of hyperbolic and inverse hyperbolic functions.

Course Outcome: At the end of the course, the students will be able to

1. Understand the concept of integration
2. Apply the knowledge of integral calculus in finding length of arc, area under curves, volume and area of surface swept by curve during revolution

Text Books

1. H. S. Hall and S. R. Knight, Higher Algebra, H. M. Publications, 1994.
2. Chandrika Prasad, Text Book on Algebra and Theory of Equations, PothishalaPvt. Ltd., 2017.
3. Richard L. Burden and J. Douglas Faires, Numerical Analysis, 9th Edition, Cengage Learning, 2012.
4. M. K. Jain, S. R. K. Iyengar and R. K. Jain, Numerical Methods for Scientific and Engineering Computation, 6th Edition, New Age International Publishers, 2012.

BSNM-23206
THEORY OF EQUATIONS

L	T	P	C
3	0	0	3

Course Objectives

Students will be able to demonstrate relation between roots and coefficients. Transform the equation through roots multiplied by a given number, increase the roots, decrease the roots, removal of terms solve the reciprocal equations. They will understand the concept of computer arithmetic and errors also find solution of nonlinear equations with different methods.

I. (9)

Euclid's algorithm, synthetic division, roots and their multiplicity. Complex roots of real polynomials occur in conjugate pairs with same multiplicity. Relation between roots and coefficients. Transformation of equations. Descartes' Rule of Signs.

II. (9)

Solution of cubic and bi-quadratic equations, Cardano's method of solving a cubic, discriminant and nature of roots of real cubic, trigonometric solutions of a real cubic with real roots. Ferrari's method for a bi-quadratic equation.

III. (9)

Computer arithmetic and errors: Floating point representation of numbers, numbers and their accuracy, significant digits, source of errors, types of errors, errors in arithmetic operations. Numerical instability.

IV. (8)

Algorithms, convergence, solution of nonlinear equations: Bisection method, False position method, Fixed point iteration method, Newton-Raphson's method, Secant method.

Course Outcome: Upon successful completion of this course students will be able to learn

1. How to find the roots of polynomial equations and relation between roots and coefficients.
2. They will know complex roots of real polynomials occur in conjugate pairs and understand Descartes' Rule of Signs.
3. Know how to transform the equation. Understand to solve cubic and bi-quadratic equations using Cardano's, Descartes and Ferrari's method.
4. They will understand the concept of Floating point representation of numbers source of errors, types of errors
5. They will be able to find solution of nonlinear equations with different methods like Bisection method, False position method, Fixed point iteration method, Newton-Raphson's method, Secant method

Text Books

1. H. S. Hall and S. R. Knight, Higher Algebra, H. M. Publications, 1994.
2. Chandrika Prasad, Text Book on Algebra and Theory of Equations, Pothishala Pvt. Ltd., 2017.
3. Richard L. Burden and J. Douglas Faires, Numerical Analysis, 9th Edition, Cengage Learning, 2012.
4. M. K. Jain, S. R. K. Iyengar and R. K. Jain, Numerical Methods for Scientific and Engineering Computation, 6th Edition, New Age International Publisher, 2012.

BSNM- 23207
ENGLISH-II

L T P C
2 0 0 2

Course objectives

To activate and reinforce the writing, grammar, vocabulary, and punctuation skills of the students.
To develop in them vital communication skills, integral to their personal, social and professional interactions.

- I.** **(6)**
The following **short novel** to be read for enhancing vocabulary and learning sentence/speech construction: *The Strange Case of Dr. Jekyll and Mr Hyde* by Robert Louis Stevenson
- II. Grammar** **(7)**
Parts of Speech, Adjectives and its degrees, Simple, compound and complex structures, Active and passive voices, Subject-verb agreement, Punctuation, Spelling rules and formation of words.
- III. Writing Skill** **(6)**
Report writing, Letter writing: Business and official letters, notices and memorandums, Precis writing.
- IV. Language Skills** **(6)**
Comprehension, Public speaking/Oral communication, Translation (Punjabi into English), Technical words/vocabulary.

Course Outcome

1. Students will improve their reading fluency skills through extensive reading.
2. Learners will increase their reading speech and comprehension of academic articles.

Text Books

1. Robert Louis Stevenson, *The Strange Case of Dr Jekyll and Mr Hyde*, Madhuban Publications, 2005
2. Wren and Martin, *High School English Grammar and Composition*, S Chand (Indian edition), 2008.
3. A J Thomson and A V Martinet, *A Practical English Grammar*, Oxford India, 2007.
4. R V Lesikar, M E Flatley, K Rentz and N Pande, *Business Communication (Making Connections in Digital World)*, Tata McGraw Hill, 2010.
5. M Frank, *Writing as Thinking: A Guided Process Approach*, Englewood Cliffs, Prentice Hall Regents.

BSNM- 23208
Punjabi-II

L T P C
2 0 0 2

ਵਿਸ਼ਾ ਉਦੇਸ਼: ਇਹ ਵਿਸ਼ਾ ਪੜ੍ਹਾਉਣ ਦਾ ਮੂਲ ਉਦੇਸ਼ ਵਿਦਿਆਰਥੀਆਂ ਨੂੰ ਪੰਜਾਬੀ ਸਾਹਿਤ ਦੇ ਵੱਖ-ਵੱਖ ਰੂਪਾਂ (ਕਵਿਤਾ, ਕਹਾਣੀ, ਲੇਖ, ਪ੍ਰਮੁੱਖ ਸਖਸ਼ੀਅਤਾਂ ਦੀਆਂ ਜੀਵਨੀਆਂ) ਦੇ ਮਹੱਤਵ ਅਤੇ ਪੰਜਾਬੀ ਵਿਆਕਰਣ ਤੋਂ ਜਾਣੂ ਕਰਵਾਉਣਾ ਹੈ।

ਭਾਗ-ਪਹਿਲਾ (ਕਵਿਤਾ)

ਪੁਰਾਣੇ ਪੰਜਾਬ ਨੂੰ ਅਵਾਜ਼ਾਂ (ਪ੍ਰੋ. ਪੂਰਨ ਸਿੰਘ)
ਕੁਰਬਾਨੀ (ਫੀਰੋਜ਼ ਦੀਨ ਸ਼ਰਫ਼)
ਖੈਰ ਪੰਜਾਬ ਦੀ (ਫੀਰੋਜ਼ ਦੀਨ ਸ਼ਰਫ਼)
ਆਉ ਨੱਚੀਏ (ਪ੍ਰੋ. ਮੋਹਨ ਸਿੰਘ)
ਨਵਾਂ ਕੌਤਕ (ਪ੍ਰੋ. ਮੋਹਨ ਸਿੰਘ)

ਭਾਗ-ਦੂਜਾ (ਕਹਾਣੀ)

ਤੈਂ ਕੀ ਦਰਦ ਨਾ ਆਇਆ (ਕਰਤਾਰ ਸਿੰਘ ਦੁੱਗਲ)
ਧਰਤੀ ਹੇਠਲਾ ਬੋਲਦ (ਕੁਲਵੰਤ ਸਿੰਘ ਵਿਰਕ)
ਦੂਜੀ ਵਾਰ ਜੇਬ ਕੱਟੀ ਗਈ (ਨਵਤੇਜ ਸਿੰਘ)

ਭਾਗ-ਤੀਜਾ (ਲੇਖ)

ਵਤਨ ਦਾ ਪਿਆਰ (ਪ੍ਰੋ. ਪੂਰਨ ਸਿੰਘ)
ਗੁਰੂ ਅਰਜਨ ਦੇਵ ਜੀ ਦੀ ਸ਼ਹਾਦਤ (ਡਾ. ਗੰਡਾ ਸਿੰਘ)
ਸ਼ਾਂਤੀ ਨਿਕੇਤਨ (ਸ.ਸ.ਅਮੋਲ)
ਅੱਥਰੂ (ਬਲਰਾਜ ਸਾਹਨੀ)

ਭਾਗ-ਚੌਥਾ (ਪ੍ਰਮੁੱਖ ਸਖਸ਼ੀਅਤਾਂ)

ਮੀਰਾ ਬਾਈ (ਕਪੂਰ ਸਿੰਘ)
ਗੁਰੂ ਰਾਮਦਾਸ ਜੀ (ਪ੍ਰਿੰ. ਜੋਧ ਸਿੰਘ)
ਜਰਨੈਲ ਹਰੀ ਸਿੰਘ ਨਲੂਆ (ਬਾਬਾ ਪ੍ਰੇਮ ਸਿੰਘ ਹੋਤੀ)

ਭਾਗ-ਪੰਜਵਾਂ (ਵਿਆਕਰਣ)

ਨਾਂਵ, ਪੜਨਾਂਵ, ਵਿਸ਼ੇਸ਼ਣ, ਕਿਰਿਆ, ਸਹਾਇਕ ਕਿਰਿਆ, ਕਿਰਿਆ ਵਿਸ਼ੇਸ਼ਣ, ਸਬੰਧਕ, ਯੋਜਕ, ਵਿਸਮਿਕ। ਪੈਰਾ ਰਚਨਾ, ਦਫ਼ਤਰੀ ਚਿੱਠੀ-ਪੱਤਰ। ਅਖਾਣ (ਕੋਈ ਦਸ) ਅਤੇ ਮੁਹਾਵਰੇ (ਕੋਈ ਦਸ)।

ਵਿਸ਼ਾ ਨਤੀਜਾ: ਵਿਦਿਆਰਥੀ ਵੱਖ-ਵੱਖ ਕਵਿਤਾਵਾਂ, ਕਹਾਣੀਆਂ, ਲੇਖਾਂ, ਪ੍ਰਮੁੱਖ ਸਖਸ਼ੀਅਤਾਂ ਦੀਆਂ ਜੀਵਨੀਆਂ ਦੇ ਸਾਰ ਅਤੇ ਸਿੱਖਿਆ ਰਾਹੀਂ ਪੰਜਾਬ ਦੇ ਇਤਿਹਾਸ ਅਤੇ ਪੰਜਾਬੀ ਸੱਭਿਆਚਾਰ ਵਾਰੇ ਜਾਣਕਾਰੀ ਪ੍ਰਾਪਤ ਕਰਕੇ ਵਧੀਆ ਜੀਵਨ ਜਾਚ ਗ੍ਰਹਿਣ ਕਰਕੇ ਪੰਜਾਬ ਪ੍ਰਾਂਤ ਅਤੇ ਭਾਰਤ ਦੇਸ਼ ਦੀ ਤਰੱਕੀ ਵਿੱਚ ਵਡਮੁੱਲਾ ਯੋਗਦਾਨ ਪਾਉਣ ਦੇ ਕਾਬਲ ਹੋ ਜਾਣਗੇ। ਪੰਜਾਬੀ ਵਿਆਕਰਣ ਦੀ ਜਾਣਕਾਰੀ ਵਿਦਿਆਰਥੀਆਂ ਨੂੰ ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਪੜ੍ਹਨ, ਲਿਖਣ ਅਤੇ ਸਮਝਣ ਵਿੱਚ ਨਿਪੁੰਨਤਾ ਪ੍ਰਦਾਨ ਕਰੇਗੀ।

ਪੁਸਤਕਾਂ ਦੀ ਸੂਚੀ

- ਦੇ ਰੰਗ (ਸੰਪਾਦਕ-ਹਰਜਿੰਦਰ ਸਿੰਘ ਢਿੱਲੋਂ ਅਤੇ ਪ੍ਰੀਤਮ ਸਿੰਘ ਸਰਗੋਧਾ, ਪੰਜਾਬੀ ਅਤੇ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ ਵਿਭਾਗ, ਗੁਰੂ ਨਾਨਕ ਦੇਵ ਯੂਨੀਵਰਸਿਟੀ ਅੰਮ੍ਰਿਤਸਰ)।
- ਆਧੁਨਿਕ ਪੰਜਾਬੀ ਵਾਰਤਕ (ਸੰਪਾਦਕ-ਡਾ. ਗੁਰਬਚਨ ਸਿੰਘ ਤਾਲਿਬ, ਪੰਜਾਬੀ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ ਅੰਮ੍ਰਿਤਸਰ)।
- ਸੰਸਾਰ ਦੀਆਂ ਪ੍ਰਸਿੱਧ ਹਸਤੀਆਂ (ਪ੍ਰਿੰ. ਤੇਜਾ ਸਿੰਘ ਅਤੇ ਪ੍ਰੋ. ਹਰਨਾਮ ਸਿੰਘ ਸ਼ਾਨ, ਪੰਜਾਬੀ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ ਅੰਮ੍ਰਿਤਸਰ)।

FOR BATCH 2023 AND ONWARDS
SARDAR BEANT SINGH STATE UNIVERSITY GURDASPUR

4. ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਬੋਧ (ਕਸਤੂਰੀ ਲਾਲ ਐਂਡ ਸੰਨਜ਼, ਅੰਮ੍ਰਿਤਸਰ)।
5. ਪੰਜਾਬੀ ਵਿਆਕਰਣ: ਸਿਧਾਂਤ ਤੇ ਵਿਹਾਰ (ਚੇਤਨਾ ਪ੍ਰਕਾਸ਼ਨ, ਲੁਧਿਆਣਾ)।

BSNM- 23209
CHEMISTRY LAB -II

L	T	P	C
0	0	4	2

Course Objectives: To provide students practical knowledge and skills about various topics taught in theory class of physical chemistry, which in turn will enhance their problem solving and analytical skills.

Crystallization:

Concept of indication of crystallization. Phthalic acid from hot water (using fluted filter paper & stem less funnel)

Acetanilide from boiling water.

Naphthalene from Ethanol

Benzoic acid from water

Physical Chemistry:

1. To determine the specific reaction rate of hydrolysis of ethyl acetate catalysed by Hydrogen ions at room temperature.
2. To study the effect of acid strength on hydrolysis of an ester.

Viscosity, Surface Tension (Pure Liquids)

3. To study the viscosity and surface tension of CCI glycerine solution in water.
4. To determine the solubility of benzoic acid at different temperatures and to determine ΔH of the dissolution process.
5. To determine the enthalpy of neutralisation of a weak acid/weak base versus strong base/strong acid and determine the enthalpy of ionisation of the weak acid/weak base.
6. To determine the enthalpy of dissolution of solid calcium chloride and calculate the lattice energy of calcium chloride from its enthalpy data using Born Haber cycle.

Text Books

1. Practical Organic Chemistry by F.G. Mann and B.C. Saunders
2. Advanced Practical Physical Chemistry by J.B. Jadav

BSNM- 23210
Physics Lab-II

L	T	P	C
0	0	4	2

Course Objectives: To learn the usage of electrical systems for the various measurements. Apply the analytical techniques and graphical analysis to the experimental data.

Students are required to perform at least 05 experiments from the following:

1. To use a Multimeter for measuring (a) Resistances, (b) AC and DC Voltages, (c) DC Current, and (d) checking electrical fuses.
2. To compare capacitances using De`Sauty's bridge.
3. Measurement of field strength B and its variation in a Solenoid (Determine dB/dx).
4. To study the Characteristics of a Series RC Circuit.
5. To study the series and parallel LCR circuit and determine its (a) Resonant frequency, (b) Quality Factor Q.
6. To determine a Low Resistance by Carey Foster's Bridge.
7. To verify the Thevenin and Norton theorem.
8. To verify the Superposition, and Maximum Power Transfer Theorem
9. To determine unknown capacitance by flashing and quenching method.
10. To study B-H curve for a ferromagnetic material using CRO.
11. To find out the frequency of AC mains using electric-vibrator.
12. To find out polarizability of a dielectric substance.
13. To determine the value of self-inductance by Maxwell Inductance/Capacitance Bridge.
14. To determine the mutual inductance of two coils.
15. To find out the horizontal component of earth's magnetic field (B_h).
16. Ballistic Galvanometer: (i) Measurement of charge and Current sensitivity (ii) Measurement of CDR (iii) Determine a high resistance by Leakage Method (iv) To determine Self Inductance of a Coil by Rayleigh's Method.

Text Books

1. Advanced Practical Physics for students, B. L. Flint and H.T. Worsnop, 1971, Asia Publishing House.
2. Advanced level Physics Practicals, Michael Nelson and Jon M. Ogborn, 4th Edition, reprinted 1985, Heinemann Educational Publishers
3. Engineering Practical Physics, S. Panigrahi & B. Mallick, 2015, Cengage Learning India Pvt. Ltd.
4. Practical Physics, G.L. Squires, 2015, 4th Edition, Cambridge University Press.
5. A Text Book of Practical Physics, I. Prakash & Ramakrishna, 11th Edn, 201, Kitab Mahal.
6. B Sc. Practical Physics, C. L. Arora, S. Chand & Co.
7. Text Book of Practical Physics, Indu Prakash and Ramakrishna, 11th Edition, 2011, Kitab Mahal, New Delhi

BSNM- 23211
Computer Algebra system: MATLAB

L	T	P	C
0	0	2	1

Course Objective: This course is designed to introduce a Computer Algebra System: MATLAB which is currently used in scientific computations. The main focus will be on introduction to basic concepts of MATLAB using simple examples

I. The MATLAB environment, scalars, variables, arrays, mathematical operations with arrays, built-in and user defined functions, graphics: two-dimensional and three-dimensional, m-files: script and functionfiles, functions: input; disp and fprintf, relational and logical operators.

II. Symbolic math: symbolic objects and expressions; collect; expand; factor; simplify; solve; diff and int commands, Programming: if-end structure; if-else-end structure; loops: for-end and while-end.

Course Outcomes After completion of the course, the students will be able to

- Visualize functions in 2-D and 3-D.
- Use symbolic tools of MATLAB for solving problems arising in various fields of applications.
- Make their own computer programs for solving problems of their interest.

Text Books

1. D. J. Higham and N. J. Higham, MATLAB Guide, 2nd Edition, Society for Industrial and Applied Mathematics (SIAM), 2005.
2. Amos Gilat, MATLAB: An Introduction with Applications, 5th Edition, John Wiley & Sons, 2014.+

BSNM-23301
ORGANIC CHEMISTRY-II

L	T	P	C
3	0	0	3

Course Objective

To provide a complete knowledge of nomenclature, structure and bonding methods of preparation and chemical reactions of the compounds related to functional groups like alcohols, phenols, aldehydes and ketones. To provide the knowledge of alkyl and aryl halides, arenes and aromaticity in detailed manner.

I. Alkyl and Aryl Halides (8)

Nomenclature and classes of alkyl halides, Chemical reactions. Mechanisms of nucleophilic substitution reaction of alkyl halides, SN2 and SN1 reactions with energy profile diagrams. Nuclear and side chain reactions. The addition-elimination and the elimination-addition mechanisms of nucleophilic aromatic substitution reactions. Relative reactivities of alkyl halides vs allyl, vinyl and aryl halides.

II. Arenes and Aromaticity (9)

Nomenclature of benzene derivatives. The aryl group. Aromatic nucleus and side chain. Structure of benzene: Molecular formula and Kekule structure. Stability and carbon carbon bond lengths of benzene, resonance structure, MO picture. Aromaticity : the Huckel's rule, aromatic ions. Aromatic electrophilic substitution-general pattern of the mechanism, role of σ and π complexes. Mechanism of nitration, halogenation, sulphonation, mercuration and Friedel Crafts reaction. Energy profile diagrams. Activating and deactivating substituents, orientation and ortho/para ratio. Side chain reactions of benzene derivatives. Methods of formation and chemical reactions of alkylbenzenes.

III. Alcohols (9)

Classification and nomenclature. Monohydric alcohols-nomenclature. Acidic nature. Reactions of alcohols. Dihydric alcohols-nomenclature, methods of formation, chemical reactions of vicinal glycols, oxidative cleavage $[\text{Pb}(\text{OAc})_4]$ and $[\text{HIO}_4]$ and pinacol-pinacolone rearrangement.

Phenols

Nomenclature, structure and bonding, Preparation of phenols, physical properties and acidic character, Comparative acidic strengths of alcohols and phenols, resonance stabilization of phenoxide ion. Reactions of phenols-electrophilic aromatic substitution, acylation and carboxylation. Mechanisms of Fries rearrangement, Claisen rearrangement, Gatterman synthesis, Reimer Tiemann reaction.

IV. Aldehydes and Ketones

(9)

Nomenclature and structure of the carbonyl group. Synthesis of aldehydes and ketones with particular reference to the synthesis of aldehydes from acid chlorides, synthesis of aldehydes and ketones using 1,3-dithianes, synthesis of ketones from nitriles and from carboxylic acids. Physical properties. Mechanism of nucleophilic additions to carbonyl group with particular emphasis on benzoin, aldol, Perkin and Knoevenagel condensations. Condensation with ammonia and its derivatives. Wittig reaction. Mannich reaction. Use of acetals as protecting group. Oxidation of aldehydes, Baeyer-Villiger oxidation of Ketones, Cannizzaro reaction. MPV, Clemmensen, Wolff-Kishner, LiAlH₄ and NaBH₄ reductions. Halogenation of enolizable ketones. Halogenation of enolizable ketones.

Course Outcome

1. This course will equip the students with the necessary chemical knowledge concerning the organic chemistry of functional groups like alcohols, phenols, aldehydes and ketones.
2. To develop skills to interpret and explain the mechanism of organic reactions involving different functional groups.
3. The students will be able to know about the complete knowledge of alkyl and aryl halides, arenes and aromaticity.

Text Books

1. Organic Chemistry, Morrison and Boyd, Prentice-Hall.
2. Fundamentals of Organic Chemistry, Solomons, John Wiley.
3. Organic Chemistry. F.A. Carey, McGraw Hill, Inc.
4. Organic Chemistry, L.G. Wade Jr. Prentice Hall.
5. Organic Chemistry Vol. I, II & III, S.M. Mukherji, S.P. Singh and R.P. Kapoor, Wiley Eastern Ltd (New Age International).

BSNM- 23302
PHYSICAL CHEMISTRY-II

L	T	P	C
3	0	0	3

Course Objective

This course will equip students with the necessary chemical knowledge concerning the fundamentals in the basic areas of physical chemistry viz. different states of matter. The problem solving skills of students are expected to be enhanced through numerical problems.

I. Thermodynamics-I (9)

Definition of thermodynamic terms: System, surroundings etc. Types of systems, intensive and extensive properties. State and path functions and their differentials. Thermodynamic process. Concept of heat and work.

First Law of Thermodynamics: Statement, definition of internal energy and enthalpy. Heat capacity, heat capacities at constant volume and pressure and their relationship. Joule's law- Joule-Thomson coefficient and inversion temperature, Calculation of w , q , dU & dH for the expansion of ideal gases under isothermal and adiabatic conditions for reversible process.

II. Thermodynamics-II (9)

Thermochemistry: Standard state, standard enthalpy of formation-Hess's Law of heat summation and its applications. Heat of reaction at constant pressure and at constant volume. Enthalpy of neutralization. Bond dissociation energy and its calculation from thermo-chemical data, temperature dependence of enthalpy. Kirchhoff's equation.

Second Law of Thermodynamics: Need for the law, different statements of the law, Carnot cycle and its efficiency, Carnot theorem. Thermodynamic scale of temperature.

Concept of Entropy : Entropy as a state function, entropy as a function of V & T , entropy as a function of P & T , entropy change in physical change, Clausius inequality, entropy as a criteria of spontaneity and equilibrium. Entropy change in ideal gases and mixing of gases.

III. Thermodynamics-III (9)

Third Law of Thermodynamics: Nernst heat theorem, statement and concept of residual entropy, evaluation of absolute entropy from heat capacity data. Gibbs and Helmholtz functions; Gibbs function (G) and Helmholtz function (A) as thermodynamic quantities, A & G as criteria for thermodynamic equilibrium and spontaneity, their advantage over entropy change, Variation of G and A with P , V and T .

Equilibrium

Chemical Equilibrium Equilibrium constant and free energy. Thermodynamic derivation of law of mass action. Determination of K_p , K_c , K_a and their relationship, Clausius-Clapeyron equation, applications. thermodynamic derivation and applications.

IV. Thermodynamics-III (9)

Third Law of Thermodynamics: Nernst heat theorem, statement and concept of residual entropy, evaluation of absolute entropy from heat capacity data. Gibbs and Helmholtz functions; Gibbs function (G) and Helmholtz function (A) as thermodynamic quantities, A & G as criteria for thermodynamic equilibrium and spontaneity, their advantage over entropy change, Variation of G and A with P, V and T.

Equilibrium

Chemical Equilibrium Equilibrium constant and free energy. Thermodynamic derivation of law of mass action. Determination of K_p , K_c , K_a and their relationship, Clausius-Clapeyron equation, applications. thermodynamic derivation and applications.

V. Introduction to Phase Equilibrium (8)

Statement and meaning of the terms-phase, component and degree of freedom, derivation of Gibbs phase rule, phase equilibria of one component system-water, CO_2 and S systems. Phase equilibria of two component systems-solid-liquid equilibria, simple eutectic-Bi-Cd, Pb-Ag systems, desilverisation of lead. Solid solutions-compound formation with congruent melting point (Mg-Zn) and incongruent melting point, ($\text{NaCl-H}_2\text{O}$), ($\text{FeCl}_3\text{-H}_2\text{O}$) and $\text{CuSO}_4\text{-H}_2\text{O}$ system. Freezing mixtures, acetone-dry ice. Non-ideal system-azeotropes-HCl-H₂O and ethanol water system. Partially miscible liquids Phenol-water, trines-thylamin-water, Nicotine-water System. Lower and upper consolute temperature, Effect of impurity on consolute temperature, immiscible liquids, steam distillation. Nernst distribution law-thermodynamic derivation and applications.

Course Outcome

1. Student should be able to Know the basic concepts and laws of thermodynamic-I,II & III.
2. They should be able to derive the various expressions of thermodynamic quantities like Entropy, Gibbs free energy, helmholtz function and their variation with P,V & T.
3. They should understand the phenomenon of chemical equilibrium and Clausius-Clapeyron equation.
4. They should familiar with concept of Phase rule, one component and two component system, nernst distribution law of thermodynamics.

Text Books

1. Thermodynamics for Chemists, S. Glasstone.
2. Chemical thermodynamics, P.A. Rock.
3. Principles of Physical Chemistry, S.H. Maron & C.F. Prutton.
4. Physical Chemistry, P.W. Atkins.
5. Physical Chemistry, Vol.2, K.L. Kapoor.
6. Physical Chemistry, K.J. Laidler.

BSNM303- 23303
OPTICS

L T P C
3 0 0 3

Course Objective

This course introduces students to the fundamental concepts of optics. Optics is concerned with the genesis and propagation of light, the changes that it undergoes and produces, and other phenomena closely associated with it. Students will also be given exposure to the production of laser light.

- I. Interference (9)**
Definition and properties of wave front, Temporal and Spatial Coherence, Young's double slit experiment, Lloyd's single mirror and Fresnel's Biprism. Phase change on reflection, Interference in Thin Films: parallel and wedge-shaped films, Newton's Rings: Measurement of wavelength and refractive index, Interferometer: Michelson Interferometer.
- II. Diffraction (9)**
Huygens Principle, Fraunhofer diffraction: Single slit. Circular aperture, Rayleigh criterion of resolution, Resolving Power of a telescope, Double slit, Multiple slits, Diffraction grating, Resolving power of grating, Fresnel diffraction pattern of a straight edge and circular aperture.
- III. Polarization (9)**
Plane polarized light, Representation of Unpolarized and Polarized light, Polarization by Reflection, Brewster's law, Malus Law, Polarization by Selective absorption by Crystals, Polarization by Scattering, Polarization by Double Refraction, Nicol Prism
- IV. Laser and Application (8)**
Lasers, Spontaneous emission, Stimulated absorption, Stimulated emission, Einstein coefficients, Einstein relations, Conditions for Laser actions, Population inversion, Different types of Laser Pumping mechanism: Optical Pumping, Electric Discharge and Electrical pumping, Resonators, Two, Three and Four level laser systems, Ruby laser, He-Negas Laser, CO₂ laser, applications of laser: Holography.

Course Outcome

1. The students will be able to learn interference, diffraction, reflection and refraction, and concept of zone plates.
2. The students will be able to know polarization of light and eventually they will learn the difference between Linearly, Circularly and Elliptically polarized light.
3. The students will also be able to know about coherent source of light, laser light and its applications.

Text Books

1. Optics: A.K. Ghatak (Tata-McGraw Hill), 1992.
2. Fundamentals of Optics: F.A. Jenkins and H.E. White (McGraw Hill), 1981.
3. Introduction to Modern Optics (2nd ed.), G.R. Fowles, Dover, ISBN 0-486-65957-7,2012.
4. Fundamentals of Optics, F.A. Jenkins & H.E. White, McGraw-Hill, 2011.
5. Schaum's Outline of Theory and Problems of Optics, E. Hecht, McGraw-Hill, ISBN 0-07-027730-3,1998.

BSNM- 23304
THERMAL PHYSICS

L T P C
3 0 0 3

Course Objectives

This course discusses relationship between the macroscopic properties of physical systems in equilibrium. The primary goal is to understand the fundamental laws of thermodynamics and their applications to various systems and processes like liquification of gases, Heat engines and refrigerators. In addition, it will also give exposure to students about the Kinetic theory of gases and transport phenomena involved.

- 1. Basic Concepts and Zeroth Law of Thermodynamics (7)**
Thermodynamic System, Boundaries, Extensive and intensive thermodynamic variables, Thermodynamic Equilibrium, Reversible, Irreversible and Quasi-static Processes, Representation of a Process on an Indicator Diagram. Zeroth Law of Thermodynamics & Concept of Temperature. Equation of state.
- 2. Heat and First Law of Thermodynamics (5)**
Concept of heat, Heat and Work: path functions, Internal energy, First Law of Thermodynamics, its differential form. Heat capacity, Work Done during Isothermal and Adiabatic Processes.
- 3. Second and third Law of Thermodynamics (15)**
Carnot's Cycle: Conversion of Heat into Work. Carnot's Cycle, Carnot engine & efficiency. Carnot Cycle as Refrigerator & coefficient of performance.
Second Law of Thermodynamics: Kelvin-Planck and Clausius Statements and their Equivalence. Carnot's Theorem. Thermodynamic Scale of Temperature and its Equivalence to Perfect Gas Scale.
Entropy: Physical Concept of Entropy, Clausius Theorem. Second Law of Thermodynamics in terms of Entropy. Principle of Increase of Entropy. Entropy Changes in Reversible and Irreversible processes. Representation of Carnot Cycle on Entropy-Temperature Diagram. Third Law of Thermodynamics, Unattainability of Absolute Zero.
- 4. Maxwell's thermodynamic relations, potentials and Applications (10)**
Thermodynamic Potentials, derivation of Maxwell's thermodynamic Relations. Applications of Maxwell's relations: cooling produced by adiabatic stretching and compression, stretching of films, change of internal energy with volume, TdS-Equations, Clausius - Clapeyron Equation, Joule-Thomson Effect, temperature inversion, liquefaction of gases. Cooling due to adiabatic demagnetization (qualitative).
- 5. Kinetic Theory of Gases (8)**
Introduction to Kinetic theory of matter, kinetic theory of perfect gases, Maxwell- Boltzmann distribution law and its experimental verification, expressions for Average speed, mean square and most probable speed. Law of equipartition of energy, degree of freedom and heat capacity of gases.

Course Outcomes (COs):

Students will gain knowledge about the basic concepts of thermodynamics, laws of thermodynamics,

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the concept of entropy, the thermodynamic potentials and their physical interpretations. Heat and energy, transformation process. Latent heat, specific heat, liquifaction of gases and working process of heat engines. Basic aspects of kinetic theory of gases, Maxwell-Boltzman distribution law, equitation etc.

Text Books

1. Statistical Physics and Thermodynamics-V.S. Bhatia, Punjab University, Chandigarh, 1977.
2. Thermal Physics, S. Garg, R. Bansal and C. Ghosh, 1993, Tata McGraw-Hill.
3. Heat and Thermodynamics, M.W. Zemasky and R. Dittman, 1981, McGraw Hill
4. Thermal Physics, A. Kumar and S.P. Taneja, 2014, R. Chand Publications.

BSNM-23305
ANALYSIS-I

L	T	P	C
3	0	0	3

Course Objective

The objective of the course on Analysis-I is to aware the students with the real line, its properties. The various concepts of sequence, infinite series. Also, students will be introduced to various tests to discuss the convergence, divergence of sequences and infinite series.

I. (9)
Series of non-negative terms, P-test, comparison tests, Cauchy's integral test, Cauchy's root test, D'Alembert ratio test, Raabe's test, De Morgan and Bertrand's test, Gauss' test, logarithmic test, Alternating series, absolute and conditional convergence, rearrangement of absolutely convergent series.

II. (10)
Riemann integral, integrability of continuous and monotonic functions, properties of integrable functions, the fundamental theorem of integral calculus, mean value theorems of integral calculus.

III. (6)
Improper integral and their convergence, comparison tests, absolute and conditional convergence, Abel's and Dirichlet's test.

IV. (10)
Beta and Gamma functions, properties of Gamma function, transformation of Gamma function, symmetrical property of Beta function, transformation of Beta function, relation between Beta and Gamma functions.

Course Outcome: At the end of the course, the student will be able to

1. Be acquainted with knowledge of convergent and divergent sequences
2. Apply the learnt tests in establishing convergence, divergence, absolute convergence and conditional convergence of infinite series.

Text Books

1. Shanti Narayan and M. D. Raisinghania, Elements of Real Analysis, S. Chand
2. Robert Wrede and Murray R. Spiegel, Advanced Calculus, 3rd Edition, Schaum's Outline Series (McGraw Hill).
3. S. Lang, Undergraduate Analysis, Springer-Verlag, New York.
4. S C Malik and Savita Arora, Mathematical Analysis, New Age International Publishers.

BSNM-23306
DIFFERENTIAL EQUATIONS

L	T	P	C
3	0	0	3

Course Objectives

The objective of this course is to introduce ordinary differential equations and basic theory of existence and uniqueness of solutions. This course further explains the analytic techniques in computing the solutions of various ordinary differential equations appearing in various fields of science and technology.

I. (10)
Exact differential equations, first order and higher degree equations solvable for x , y and $p=dy/dx$. Clairaut's form, singular solution as an envelope of general solutions. Geometric meaning of a differential equation. Orthogonal trajectories. Linear differential equations with constant coefficients.

II. (10)
Linear differential equations with variable coefficients: Cauchy and Legendre equations. Linear differential equations of second order- transformation of the equation by changing the dependent variable/ the independent variable, methods of variation of parameters and reduction of order, Simultaneously differential equations.

III. (7)
Partial differential equation: Formation of first and second order equations, linear equation of first order, integral surfaces passing through a given curve, surfaces orthogonal to a given system of surfaces.

IV. (8)
Nonlinear first order partial differential equations: Charpit's method, Higher order linear partial differential equations with constant coefficients: complementary function, particular integral.

Course Outcome: At the end of the course, the students will be able to

1. Understand the basic definitions to know about ordinary differential equations, its various types and their solutions
2. Visualize the geometrical meaning of first order differential equation.

Text Books

1. W E Boyce and R C DiPrima, Elementary Differential Equations and Boundary Value Problems, 9th Edition, Wiley
2. R K Jain and S R K Iyengar, Advanced Engineering Mathematics, 4th Edition, Narosa Publishing House Pvt Ltd, New Delhi

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3. I N Sneddon, Elements of Partial Differential Equations, McGraw-Hill
4. S L Ross, Differential Equations, John Wiley & Sons
5. M D Raisinghania, Advanced Differential Equations, 19th Edition, S. Chand

BSNM-23307
STATICS AND DYNAMICS

L	T	P	C
3	0	0	3

Course Objective

To develop the ability to apply Newtonian mechanics to model and predict the responses of simple dynamical system (particle and rigid body) subject to applied forces. Learn to solve equilibrium of rigid bodies including the calculations of moment of force, inertia moments of solid bodies and basic structural analysis and be able to determine the requirement for the equilibrium of particles and solid bodies.

I. (9)
Basic notions. Composition and resolution of concurrent forces-parallelogram law of forces, Components of force in given directions, resolved parts of a force, resultant of any number of coplanar concurrent forces.

II. (8)
Equilibrium conditions for coplanar concurrent forces, equilibrium of a body resting on a smooth inclined plane, equilibrium of three forces acting at a point, triangle law of forces, theorem, Lami's theorem, parallel forces.

III. (10)
Motion of a particle with constant acceleration, acceleration of falling bodies, motion under gravity, motion of a body projected vertically upwards: Newton's Laws of Motion, Motion of two particles connected by a string, motion along a smooth inclined plane, constrained motion along a smooth inclined plane. Variable acceleration: Simple harmonic motion, elastic string.

IV. (8)
Curvilinear motion of a particle in a plane: Definition of velocity and acceleration, projectiles, motion in a circle. Work, power, conservative fields and the potential energy, work done against gravity, potential energy of a gravitational field.

Course Outcome: At the end of this course, the student will be able to demonstrate ability in the following areas:

1. Determine the resultant of force systems.
2. Calculate position, displacement, velocity and acceleration of a particle
3. Use principles of work and energy for basic dynamic analysis
4. Distinguish between particles and rigid bodies and their analysis

Text Books

1. S. L. Loney, Statics, Macmillian and Company London.

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2. R. S. Verma, A Textbook on Statics, Pothishala Pvt. Ltd. Allahabad.
3. S. L. Loney, An Elementary Treatise on the Dynamics of a Particle and of
4. Rigid bodies, Cambridge University Press, 1956.
5. M. Ray, A Textbook on Dynamics, S. Chand & Company, 1989.

BSNM-23308
ENGLISH-III

L T P C
2 0 0 2

Course Objective

Providing the student with information about studies of vocabulary based on a quantitative analysis of what people actually use in speech and writing.

I.

(6)

Textbook entitled 'Prism: Spoken and Written Communication, Prose & Poetry' published by Orient Longman

For enhancing vocabulary and learning sentence/speech construction:

Prose:

- 1) The Bet – Anton Chekov
- 2) An Astrologer's Day – R. K. Narayan
- 3) The Gift of the Magi – O' Henry Poetry:
 - 1) The Felling of the Banyan Tree – Dilip Chitre
 - 2) Stay Calm – Grenville Kleiser

II. Grammar and Vocabulary

(6)

Modal auxiliaries, Gerunds Infinitives; Participles; Usage of Conjunctions; Scientific & Technical Vocabulary;

III. Reading & Writing Skills

(6)

Note Making and Note Taking; Writing abstracts & summaries

IV. Spoken Skills

(7)

- 1) Meeting People, Exchanging Greetings and Taking Leave
- 2) Introducing Yourself
- 3) Introducing People to Others
- 4) Answering the Telephone and Asking for Someone
- 5) Dealing with a Wrong Number
- 6) Taking and Leaving Messages
- 7) Making Inquiries on the Phone
- 8) Calling for Help in an Emergency

Course Outcome

1. To take notes-short and fast-on specific information.
2. Be able to speak communicatively.
3. Students will apply learned grammatical and idiomatic structures to self-monitor speech

Text Books

FOR BATCH 2023 AND ONWARDS
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1. William Zinsser. On Writing Well. Harper Resource Book. 2001
2. Robert Louis Stevenson, The Strange Case of Dr Jekyll and Mr Hyde, Madhuban Publications, 2005
3. Wren and Martin, High School English Grammar and Composition, S Chand (Indian edition), 2008.
4. A J Thomson and A V Martinet, A Practical English Grammar, Oxford India, 2007
5. R V Lesikar, M E Flatley, K Rentz and N Pande, Business Communication (Making Connections in Digital World), Tata McGraw Hill, 2010
6. M Frank, Writing as Thinking: A Guided Process Approach, Englewood Cliffs, Prentice Hall Regents.

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BSNM-23309
PUNJABI-III

L T P C
2 0 0 2

Unit	Contents	Contact Hours
I	<p>ਕਵਿਤਾ ਭਾਗ:</p> <p>ਭਾਈ ਵੀਰ ਸਿੰਘ: ਸਮਾਂ, ਚੜ੍ਹਮਾ</p> <p>ਪ੍ਰੋ. ਪੂਰਨ ਸਿੰਘ : ਪੰਜਾਬ ਨੂੰ ਕੂਕਾਂ ਸੌ, ਹੱਲ ਵਾਹੁਣ ਵਾਲੇ</p> <p>ਪ੍ਰੋ.ਮੋਹਨ ਸਿੰਘ : ਮਾਂ, ਕੋਈ ਆਇਆ ਸਾਡੇ ਵਿਹੜੇ, ਪਿਆਰ ਪੰਥ</p> <p>ਅੰਮ੍ਰਿਤਾ ਪ੍ਰੀਤਮ: ਆਖਾਂ ਵਾਰਿਸ ਸ਼ਾਹ ਨੂੰ, ਅੰਨਦਾਤਾ</p>	12
II	<p>ਕਹਾਣੀ ਭਾਗ:</p> <p>ਸੰਤ ਸਿੰਘ ਸੇਖੋਂ : ਪੇਮੀ ਦੇ ਨਿਆਣੇ</p> <p>ਸੁਜਾਨ ਸਿੰਘ : ਕੁਲਫੀ</p> <p>ਕੁਲਵੰਤ ਸਿੰਘ ਵਿਰਕ : ਤੂੜੀ ਦੀ ਪੰਡ</p> <p>ਗੁਰਦਿਆਲ ਸਿੰਘ : ਸਾਂਝ</p>	11
III	<p>ਸਵਰ ਤੇ ਵਿਅੰਜਨ ਧੁਨੀਆਂ ਦਾ ਨਿਖੇੜਾ ਤੇ ਵਰਗੀਕਰਨ ਦੁੱਤ ਵਿਅੰਜਨ ਤੇ ਸੰਯੁਕਤ ਵਿਅੰਜਨ ਅਗੇਤਰ, ਪਿਛੇਤਰ</p>	12

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IK Gujral Punjab Technical University Jalandhar
B.Sc. (Non-Medical) Batch 2018 onwards

IV	ਪੰਜਾਬੀ ਦੀਆਂ ਧੁਨੀਆਂ ਦੇ ਪਰਿਵਰਤਨ ਦੀਆਂ ਦਿਸ਼ਾਵਾਂ : ਲੋਪ, ਆਗਮ, ਵਿਕਾਰ, ਵਿਸ਼ਮੀਕਰਨ, ਵਿਪਰਜ। ਪੰਜਾਬੀ ਵਾਕ ਬਣਤਰ ਦਾ ਵਿਸਤਾਰ ਪੂਰਵਕ ਅਧਿਐਨ	10
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Reference Books

S.No.	Author(s)	Title of the Book	Publisher/Year
1	ਡਾ. ਮਹਿਲ ਸਿੰਘ (ਸੰਪ.)	ਸਾਹਿਤ ਦੇ ਰੰਗ	ਰਵੀ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ, ਅੰਮ੍ਰਿਤਸਰ।
2	ਡਾ. ਸੁਖਵਿੰਦਰ ਸਿੰਘ ਸੰਘਾ	ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਵਿਗਿਆਨ	ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਅਕਾਦਮੀ, ਜਲੰਧਰ

BSNM-23310
CHEMISTRY LAB III

L	T	P	C
0	0	4	2

Quantitative Analysis

Volumetric Analysis

1. Determination of acetic acid in commercial vinegar using NaOH.
2. Determination of alkali content-antacid tablet using HCl.
3. Estimation of calcium content in chalk as calcium oxalate by permanganometry.
4. Estimation of hardness of water by EDTA.
5. Estimation of ferrous and ferric by dichromate method.
6. Estimation of copper using sodiumthiosulphate.

Gravimetric Analysis

Analysis of Cu as CuSCN and Ni as Ni (dimethylgloxime)

Organic Chemistry Laboratory

Techniques Thin Layer

Chromatography

Determination of R_f values and identification of organic compounds.

1. Separation of green leaf pigments (spinach leaves may be used).
2. Preparation and separation of 2, 4. dinitrophenylhydrazones of acetone, 2-butanone, 2-Butanone, hexan-2 and 3-one using toluene and light petroleum (40 : 60).
3. Separation of a mixture of dyes using cyclohexane and ethyl acetate (8.5:1.5).

Text Books

1. Practical Organic Chemistry by F.G. Mann and B.C. Saunders
2. Practical Inorganic Chemistry by J.R. Barrante G. Marr and B.W. Rockett
3. Vogel's Inorganic Quantitative Analysis

BSNM-23311
PHYSICS LAB-III

L	T	P	C
0	0	4	2

Students are required to perform at least 5 experiments from the following:

1. To study the laser beam characteristics like; wave length using diffraction grating aperture & divergence.
2. Study of diffraction using laser beam and thus to determine the grating element.
3. To study laser interference using Michelson's Interferometer.
4. To study wavelength of sodium light using Newton Rings.
5. To determine the numerical aperture of a given optic fibre and hence to find its acceptance angle.
6. To find the refractive index of a material/glass using spectrometer.
7. To find the refractive index of a liquid using spectrometer
8. To find the velocity of ultrasound in liquid.
9. To determine the specific rotation of sugar using Laurent's half-shade polarimeter.
10. To determine the coefficient of thermal conductivity of a bad conductor using Lee's disc apparatus.
11. To compare heat transfer between different material surface and the black body surface by radiation.
12. To find the emissivity of different material surface.

Text Books

1. Advanced Practical Physics for students, B. L. Flint and H.T. Worsnop, 1971, Asia Publishing House.
2. Advanced level Physics Practicals, Michael Nelson and Jon M. Ogborn, 4th Edition, reprinted 1985, Heinemann Educational Publishers
3. Engineering Practical Physics, S.Panigrahi & B. Mallick, 2015, Cengage Learning India Pvt. Ltd.
4. Practical Physics, G.L. Squires, 2015, 4th Edition, Cambridge University Press.
5. A Text Book of Practical Physics, I. Prakash & Ramakrishna, 11th Edn, 201, Kitab Mahal.
6. B Sc. Practical Physics, C. L. Arora, S. Chand & Co.
7. A Text Book of Practical Physics, Indu Prakash and Ramakrishna, 11th Edition, 2011, Kitab Mahal, New Delhi.

BSNM-23401
INORGANIC CHEMISTRY-III

L T P C
3 0 0 3

Course objective

The objective of this course is to study the chemistry of coordination complexes, lanthanides, actinides elements and behaviour of molecules in different solvents. This course also helps in understanding the chemistry of bioinorganic chemistry.

I. Coordination Compounds (9)

Werner's coordination theory and its experimental verification, effective atomic number concept, chelates, nomenclature of coordination compounds, isomerism in coordination compounds, valence bond theory of transition metal complexes.

II. Non-aqueous Solvents (10)

Physical properties of a solvent, types of solvents and their general characteristics, reactions in non-aqueous solvents with reference to liquid NH₃ and liquid SO₂.

Oxidation and Reduction

Use of redox potential data-analysis of redox cycle, redox stability in water-Frost, Latimer and Pourbaix diagrams.

III. Chemistry of Lanthanide Elements (8)

Electronic structure, oxidation states and ionic radii and lanthanide contraction. Electronic absorption and magnetic properties of lanthanides.

Chemistry of Actinides

General features and chemistry of actinides, similarities between the later actinides and the later lanthanides. Electronic and magnetic properties of actinides and their general comparison with the lanthanide elements.

IV. Bioinorganic Chemistry (8)

Essential and trace elements in biological processes, metalloporphyrins and special reference to haemoglobin and myoglobin. Biological role of alkali and alkaline earth metal ions with special reference to Ca²⁺.

Course Outcome

1. This course will equip the students with the necessary knowledge of lanthanides and actinides elements.
2. To know about some bioinorganic compounds
3. To categorise coordination compounds and understand the concept of valence bond theory.
4. To understand the properties of non-aqueous solvents and analysis of redox reactions.

Text Books

1. J.D. Lee, Concise Inorganic Chemistry, 4th Ed.
2. J.E. Huheey, Inorganic Chemistry, Harper & Row.
3. F.A.Cotton and G. Wilinon, Advanced Inorganic Chemistry, Interscience Publishers.
4. N.N. Greenwood and A. Earnshaw, Chemistry of Elements, Pergamon Press.
5. D.F.C. Shriver, P.W. Atkins and C.H. Langford, Inorganic Chemistry, ELBS Oxford, 1991s

BSNM-23402
ORGANIC CHEMISTRY-III

L T P C
3 0 0 3

Course objective

The objective of this course is to study the physical and chemical properties of organic compounds. To provide a complete knowledge of nomenclature, structure and bonding, methods of preparation and chemical reactions of the compounds related to functional groups like carboxylic acids, ethers and epoxides, organic compounds of nitrogen and heterocyclic compounds.

I. Carboxylic Acids (9)

Nomenclature, structure and bonding, physical properties, acidity of carboxylic acids, effects of substituents on acid strength. Reactions of carboxylic acids. Hell-Volhard-Zelinsky reaction. Synthesis of acid chlorides, esters and amides. Reduction of carboxylic acids. Mechanism of decarboxylation.

Carboxylic Acids Derivatives

Structure and nomenclature of acid chlorides, esters, amides and acid anhydrides, Relative stability & reactivity of acyl derivatives. Physical properties, interconversion of acid derivatives by nucleophilic acyl substitution. Preparation of carboxylic acid derivatives, chemical reactions. Mechanisms of esterification and hydrolysis (acidic and basic).

II. Ethers and Epoxides (9)

Nomenclature of ethers and methods of their formation, physical properties. Chemical reaction cleavage and autoxidation, Ziesel's method. Synthesis of epoxides. Acid and base-catalyzed ring opening of epoxides, orientation of epoxide ring opening, reactions of Grignard and organolithium reagents with epoxides.

Organometallic Compounds

Organomagnesium Compounds: The Grignard reagents-formation, structure and chemical reactions. Organolithium Compounds: Formation and chemical reactions. Organozinc and Organo copper Compounds: Nomenclature, structural features, Methods of formation and chemical reactions.

III. Organic Compounds of Nitrogen (9)

Preparation of nitroalkanes and nitroarenes. Chemical reactions of nitroalkanes, Mechanisms of nucleophilic substitution in nitroarenes and their reduction in acidic, neutral and alkaline media. Reactivity, Structure and nomenclature of amines, Methods of preparation of amines by Reductive amination of aldehydic and ketonic compounds, Gabriel-phthalimide reaction and Hofmann bromamide reaction. Physical properties. Stereochemistry of amines. separation of a mixture of primary, secondary and tertiary amines. Structural features effecting basicity of

amines. Amine salts as phase-transfer catalysts.

IV. Heterocyclic Compounds (8)

Introduction: Molecular orbital picture and aromatic characteristics of pyrrole, furan, thiophene and pyridine. Methods of synthesis and chemical reactions with particular emphasis on the mechanism of electrophilic substitution. Mechanism of nucleophilic substitution reactions in pyridine derivatives. Comparison of basicity of pyridine, piperidine and pyrrole.

Course Outcome

1. This course will equip the students with the necessary chemical knowledge concerning the organic chemistry of functional groups.
2. To develop skills to interpret and explain the mechanism of organic reactions involving different functional groups.
3. The students will be able to know about the heterocyclic compounds in detailed manner.

Text Books

1. Organic Chemistry. F.A. Carey, McGraw Hill, Inc. 8th edition.
2. Organic Chemistry, Morrison and Boyd, Prentice Hall
3. Heterocyclic Chemistry, J.A. Joule, K. Mills and G.F. Smith, 3rd edition, Indian reprint, 2004. Chennai Microprint Pvt. Ltd.
4. Heterocyclic Chemistry, T.L. Gilchrist, Longman Scientific Technical
5. Organic Chemistry Vol. I, II & III, S.M. Mukherji, S.P. Singh and R.P. Kapoor, Wiley Eastern Ltd (New Age International).
6. Introduction to organic chemistry, Stritwieser, Heathcock and Kosover, Macmillan.

BSNM-23403
Waves Vibrations

L T P C
4 0 0 4

Course Objective

The objective of the course provides an exposure about simple harmonic motions, damped harmonic motions and forced oscillations. Students learn about the different waves, propagation of waves in various mediums and reflection/transmission of waves at the interface of mediums.

I. Simple and Damped Harmonic Motion (10)

Simple harmonic motion, energy of a SHO, Compound pendulum, Torsional pendulum, Electrical Oscillations, Lattice Vibrations, Transverse Vibrations of a mass on a string, Anharmonic Oscillations. Damped simple harmonic motion, Decay of free Vibrations due to damping, types of damping, Determination of damping coefficients – Logarithmic decrement, relaxation time and Q-factor. Electromagnetic damping.

II. Forced Vibrations and Resonance (9)

Forced mechanical and electrical oscillator, Transient and Steady State Oscillations, Displacement and velocity variation with driving force frequency, Variation of phase with frequency resonance, Power supplied to forced oscillator by the driving force. Q-factor and band width of a forced oscillator, Electrical and nuclear magnetic resonances.

III. Coupled Oscillations (8)

Stiffness coupled oscillators, Normal coordinates and modes of vibrations. Inductance coupling of electrical oscillators, Normal frequencies, Forced vibrations and resonance for coupled oscillators, Masses on string-coupled oscillators.

IV. Waves in Physical Media (10)

Types of waves, wave equation (transverse) and its solution characteristics impedance of a string, Impedance matching, Reflection and Transmission of waves at boundary, Energy of vibrating string, wave and group velocity.

Course Outcome: At the end of the course, the student will be able to

1. Understand the simple and damped harmonic motion of an oscillator.
2. Understand Forced Vibrations and phenomenon of Resonance
3. Apply the Coupled oscillator to the real-life problems.
4. Understand the transmission of signals and Electromagnetic Waves.

Text Books

1. Text Book of Vibrations and Waves: S.P. Puri (Macmillan India), 2004.
2. The Physics of Vibrations and Waves: H.J. Pain (Wiley and ELBS), 1976.

**BSNM-23404
ELECTRONICS**

L T P C
3 0 0 3

Course Objective

This course will enable the students to understand and gain the knowledge on semiconductor devices. It will also acquaint the students with construction, theory and characteristics of the electronic devices.

I. P.N. Junction (10)

Intrinsic/Extrinsic semiconductor, Fermi level, Charge carries in semiconductors, PN junctions, depletion region, current components in pn junction, Characteristic of pn junction diode, pn junction as rectifier, characteristics and applications of Zener diode, Photodiode, LED and photocells.

II. Electronic Devices (10)

Bipolar junction transistor, current components in transistors, CB, CE, CC configuration, h-parameters, transistor biasing, transistor as an amplifier, Emitter follower, characteristics and applications of FET, MOSFET.

III. Transistor Circuits (10)

Feedback amplifiers; classification of amplifiers, feed-back concept, Sinusoidal oscillations; phase shift oscillators, Wien Bridge Oscillator, Crystal oscillator, Basic idea about AM modulation and demodulations, Oscilloscope.

IV. Digital Principles (10)

Number system, Decimal, binary, Octal, hexadecimal, logic gates, AND, OR, NOT, NAND, NOR, XOR, XNOR, Karnaugh map techniques.

Course Outcome: On the successful completion of the course, student will be able to

1. Explain the structure of the basic electronic devices
2. Understand the characteristics and operations of special diodes
3. Understand the characteristics and operations of FET and UJT
4. Understand the characteristics and operations of transistors

Text Books

1. Integrated Electronics: J.Millman and C.C.Halkias (Tata McGraw Hill,2001).
2. Electronic Devices & Circuits–J.Millman and C.C.Halkias (Tata McGraw Hill, 2009).
3. Digital Principles & Applications–P.Malvine & Leach (Tata McGraw Hill,1993)

BSNM-23405
ANALYSIS-II

L T P C
3 0 0 3

Course Objective

The objective of the course is to have knowledge on limit theorems on functions, limits of functions, continuity of functions and its properties, uniform continuity, differentiability of functions, algebra of functions and Taylor's theorem and, its applications.

I. Sequence of functions

(9)

pointwise and uniform convergence, Cauchy's criterion for uniform convergence, Test (Mn -test) for uniform convergence, uniform convergence and continuity, uniform convergence and integration, uniform convergence and differentiation.

II. Series of functions

(10)

Pointwise and uniform convergence, Cauchy's criterion for uniform convergence, Weierstrass's M-test test, Abel's test, Dirichlet's test, uniform convergence and continuity, uniform convergence and integration, uniform convergence and differentiation. Weierstrass approximation theorem (Statement only).

III.

(10)

Vector differentiation, Gradient, Divergence and Curl with their properties and applications. Vector Integration: Line, Surface and Volume integration. Gauss divergence theorem, Stokes' theorem, Green's theorem.

IV. Fourier series

(6)

Fourier expansion of piecewise monotonic functions, Fourier series for odd and even functions, half range series. Fourier series in the interval $[0, 2\pi]$, $[-1, 1]$ and $[a, b]$.

Course Outcome

On the completion of the course, students will have working knowledge on the concepts and theorems of the elementary calculus of functions of one real variable. They will work out problems involving derivatives of function and their applications. Students will also be able to compute the Fourier Series and its complex form.

Text Books

1. Tom Apostol, Mathematical Analysis, Narosa Publishing House, New Delhi, 1985.
2. Shanti Narayan, M. D. Raisinghania, Elements of Real Analysis, S. Chand & Company, 2018.
3. S. C. Malik, Savita Arora, Mathematical Analysis, New Age International Publishers, 2017.
4. Erwin Kreyszig, Advanced Engineering Mathematics, John Wiley & Sons Inc, New York, 1999

BSNM-23406
LINEAR ALGEBRA

L	T	P	C
4	0	0	4

Course Objective

To make the students become familiar with the basic concepts of linear algebra with a thorough understanding of vector spaces, linear transformations and matrix operations enhancing the students' ability to reason mathematically and able to apply this knowledge in the field of science and engineering.

I. (10)

Linear independence of row and column vectors, row rank, column rank and rank of a matrix and their equivalence. Applications of matrices to a system of linear equations (both homogeneous and non-homogeneous). Theorems on consistency of a system of linear equations (both homogeneous and non-homogeneous).

II. (10)

Eigenvalues, eigenvectors and characteristic equation of a matrix, Cayley-Hamilton theorem and its use in finding inverse of a matrix. Diagonalization.

III. (10)

Vector Space: Definition and Examples of Vector Spaces, Subspaces, Algebra of subspaces, Linear span, Linear dependence and independence of vectors, Basis and dimension of a vector space, Basis and dimension of subspace, Direct sums and complements.

IV. (10)

Linear transformations, Rank and Nullity of a linear transformation, Vector space of linear transformations. Linear transformations and matrices, Change of basis.

Course Outcome: Upon successful completion of this course, students will:

1. Solve systems of linear equations by matrix inversion and rank method.
2. Demonstrate understanding of the concepts of vector space and subspace.
3. Demonstrate understanding of linear independence, span, and basis.
4. Determine eigenvalues and eigenvectors and solve eigenvalue problems.
5. Apply principles of matrix algebra to linear transformations.

Text Books

1. P. B. Bhattacharya, S. K. Jain, S. R. Nagpaul, First Course in Linear Algebra, New Age International Publishers
2. Bernard Kolman, David R. Hill, Elementary Linear Algebra with Applications, Pearson
3. Vivek Sahai, Vikas Bist, Linear Algebra, Narosa, 2017.

**BSNM-23407
ENGLISH-IV**

L T P C
2 0 0 2

Course Objective

Train students to speak and communicate accurately and fluently. Developing skills relevant to listening, concentration on and fast processing of information.

- I. Textbook entitled ‘Prism: Spoken and Written Communication, Prose & Poetry’ published by Orient Longman (6)**

For enhancing vocabulary and learning sentence/speech construction:

I. Prose

- 1) Socrates and the Schoolmaster – F. L. Brayne
- 2) With the Photographer – Stephen Leacock

II. Poetry

- 1) On Television – Roald Dahl
- 2) Say Not the Struggle Naught Availeth – Arthur Hugh Clough
- 3) Abou Ben Adhem – James Leigh Hunt

- II. Grammar and Vocabulary (6)**

Transformation of sentences; Tenses; Active/Passive Voice; Narration

- III. Reading & Writing Skills (4)**

Analytical reports; Drafting of career documents: Job Applications/Resume/CV

- IV. Spoken Skills (6)**

1. Getting People’s Attention and Interrupting
2. Giving Instructions and Seeking Clarifications
3. Making Requests and Responding to Requests
4. Asking for Directions and Giving Directions
5. Thanking Someone and Responding to Thanks
6. Inviting and Accepting and Refusing an Invitation
7. Apologizing and Responding to an Apology
8. Asking for, Giving and Refusing Permission

Course Outcome

1. It helps the students to understand the creative resources of language in poetry and prose.
2. They will understand that grammar can be seen as a flexible and useful tool for meaning making.

FOR BATCH 2023 AND ONWARDS
SARDAR BEANT SINGH STATE UNIVERSITY GURDASPUR

Text Books

1. William Zinsser. *On Writing Well*. Harper Resource Book. 2001
2. Robert Louis Stevenson, *The Strange Case of Dr Jekyll and Mr Hyde*, Madhuban Publications, 2005
3. Wren and Martin, *High School English Grammar and Composition*, S Chand (Indian edition), 2008
4. A J Thomson and A V Martinet, *A Practical English Grammar*, Oxford India, 2007
5. R V Lesikar, M E Flatley, K Rentz and N Pande, *Business Communication (Making Connections in Digital World)*, Tata McGraw Hill, 2010
6. M Frank, *Writing as Thinking: A Guided Process Approach*, Englewood Cliffs, Prentice Hall Regents.

FOR BATCH 2023 AND ONWARDS
SARDAR BEANT SINGH STATE UNIVERSITY GURDASPUR

BSNM-23408
Punjabi-IV

L T P C
2 0 0 2

Course Objective

Unit	Contents	Contact Hours
I	<p>ਡਾ.ਹਰਿਭਜਨ ਸਿੰਘ:</p> <p style="text-align: center;">ਅਪ੍ਰਮਾਣਿਕ, ਤੇਰੇ ਹਜ਼ੂਰ ਮੇਰੀ ਹਾਜ਼ਰੀ ਦੀ ਦਾਸਤਾਨ</p> <p>ਸ਼ਿਵ ਕੁਮਾਰ ਬਟਾਲਵੀ:</p> <p style="text-align: center;">ਕੀਡਿਆਲੀ ਬੋਰੂ, ਧਰਮੀ ਬਾਬਲ ਪਾਪ ਕਮਾਇਆ, ਰੁੱਖ</p> <p>ਪਾਸ਼:</p> <p style="text-align: center;">ਇਨਕਾਰ,ਸਭ ਤੋਂ ਖਤਰਨਾਕ,ਦਹਿਕਦੇ ਅੰਗਿਆਰਾਂ 'ਤੇ</p> <p>ਸੁਰਜੀਤ ਪਾਤਰ:</p> <p style="text-align: center;">ਹੁਣ ਘਰਾਂ ਨੂੰ ਪਰਤਣਾ, ਕੁਝ ਕਿਹਾ ਤਾਂ..., ਪੁਲ</p>	12
II	<p>ਕਹਾਣੀ ਭਾਗ:</p> <p>ਸੰਤੋਖ ਸਿੰਘ ਧੀਰ:</p> <p style="text-align: center;">ਕੋਈ ਇਕ ਸਵਾਰ</p> <p>ਪ੍ਰੇਮ ਪ੍ਰਕਾਸ਼:</p> <p style="text-align: center;">ਲੱਛਮੀ</p> <p>ਮੋਹਨ ਭੰਡਾਰੀ :</p> <p style="text-align: center;">ਘੋਟਣਾ</p> <p>ਵਰਿਆਮ ਸਿੰਘ ਸੰਧੂ :</p> <p style="text-align: center;">ਆਪਣਾ ਆਪਣਾ ਹਿੱਸਾ</p>	11
III	<p>ਕੰਪਿਊਟਰ ਦੀ ਪਰਿਭਾਸ਼ਾ, ਡਾਟਾ ਸਟੋਰੇਜ਼ ਡਿਵਾਈਸਜ਼, ਟਾਈਪਿੰਗ ਦੀ ਮਹੱਤਤਾ, ਫਾਈਡ ਐਂਡ ਰੀਪਲੇਸ : ਫਾਈਡ ਐਂਡ ਚੇਜ਼ ਦ ਟੈਕਸਟ, ਸਪੈਲ ਚੈੱਕਰ</p>	12

FOR BATCH 2023 AND ONWARDS
SARDAR BEANT SINGH STATE UNIVERSITY GURDASPUR

B.Sc. (Nursing) Batch 2018 onwards

	ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਤੇ ਇੰਟਰਨੈੱਟ : ਈ.ਨਿਊਜਪੇਪਰ, ਵਿਕੀਪੀਡੀਆ	
IV	ਸਾਹਿਤ ਦੇ ਰੂਪ : ਕਵਿਤਾ, ਵਾਰਤਕ, ਕਹਾਣੀ, ਨਾਵਲ	10

Reference Books

S.No.	Author(s)	Title of the Book	Publisher/Year
1	ਡਾ. ਮਹਿਲ ਸਿੰਘ (ਸੰਪ.)	ਸਾਹਿਤ ਦੇ ਰੰਗ	ਰਵੀ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ, ਅੰਮ੍ਰਿਤਸਰ।
2	ਡਾ. ਸੁਖਵਿੰਦਰ ਸਿੰਘ ਸੰਘਾ	ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਵਿਗਿਆਨ	ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਅਕਾਦਮੀ, ਜਲੰਧਰ
3	ਰਤਨ ਸਿੰਘ ਜੱਗੀ	ਸਾਹਿਤ ਦੇ ਰੂਪ	ਪੰਜਾਬੀ ਯੂਨੀਵਰਸਿਟੀ, ਪਟਿਆਲਾ

Course Outcome

BSNM-23409
CHEMISTRY LAB IV

L	T	P	C
0	0	4	2

Qualitative Analysis

Detection of elements

1. Nitrogen,
2. Sulphur
3. Halogens

Detection of functional groups

1. Phenolic
2. carboxylic,
3. carbonyl,
4. esters,
5. carbohydrates,
6. amines, amides, nitro and anilide

in simple organic compounds and preparing their derivatives

Text Books

1. Practical Organic Chemistry by F.G. Mann and B.C. Saunders

FOR BATCH 2023 AND ONWARDS
SARDAR BEANT SINGH STATE UNIVERSITY GURDASPUR

BSNM-23410
Physics Lab IV

L	T	P	C
0	0	4	2

Students are required to perform at least 5 experiments from the following:

1. To determine the value of horizontal component of Earth's magnetic field B_h .
2. To determine unknown capacitance by flashing and quenching method.
3. To study the magnetic field of a circular coil carrying current.
4. To find out polarizability of a dielectric substance.
5. To determine the frequency of an electrically maintained tuning fork by i) Transverse mode of vibration ii) Longitudinal mode of vibration
6. To find out the frequency of AC mains using electric-vibrator/sonometer.
7. Experiment to study Doppler effect
8. To study V-I characteristic of a Ge-Si junction.
9. Analyze the suitability of a given Zener diode as a power regulator.
10. To study the band gap of a Ge semiconductor.
11. To study the the band gap of a Si semiconductor.

Text Books

1. Advanced Practical Physics for students, B. L. Flint and H.T. Worsnop, 1971, Asia Publishing House.
2. Advanced level Physics Practicals, Michael Nelson and Jon M. Ogborn, 4th Edition, reprinted 1985, Heinemann Educational Publishers
3. Engineering Practical Physics, S.Panigrahi & B.Mallick, 2015, Cengage Learning India Pvt. Ltd
4. Practical Physics, G.L. Squires, 2015, 4th Edition, Cambridge University Press.
5. A Text Book of Practical Physics, I.Prakash & Ramakrishna, 11th Edn, 201, KitabMahal.
6. B Sc. Practical Physics, C. L. Arora, S. Chand & Co.
7. A Text Book of Practical Physics, Indu Prakash and Ramakrishna, 11th Edition, 2011, Kitab Mahal, New Delhi

BSNM-23411
Mathematica Software

L	T	P	C
0	0	2	1

I. The structure of MATHEMATICA, notebook interfaces, constants, variables, algebraic calculations, four kinds of brackets, lists, tables, expressions, functions, built-in functions, functional operations, graphics, patterns, manipulating lists, transformation rules, evaluation of expressions, modularity, manipulating notebooks, relational and logical operators.

II. Symbolic math commands: D; Integrate; Sum; Product; Solve; Eliminate; Reduce; Series; Limit; Minimize; Programming: conditionals; loops: Do; For and While.

Text Books

1. Wolfram, S., The MATHEMATICA Book, 5th revised edition. Wolfram Media Inc, 2004.
2. Abell, M. and Braselton, J., Mathematica by Example, 5th Edition. Academic Press, 2017.

BSNM-23501
INORGANIC CHEMISTRY-IV

L T P C
3 0 0 3

Course Objective

The objective of this course is to study the various aspects of transition metal complexes and organometallic compounds in detail. This will help the student to apply the knowledge in the advanced field of inorganic chemistry.

I. Metal-ligand Bonding in Transition Metal Complexes (9)

Valence bond theory, Limitations of valence bond theory, an elementary idea of crystal-field theory, crystal field splitting in octahedral, tetrahedral and square planar complexes, factors affecting the crystal-field parameters.

II. Magnetic Properties of Transition Metal Complexes (9)

Types of magnetic behaviour, methods of determining magnetic susceptibility, spin-only formula. L-S coupling, correlation of μ_s and μ_{eff} values, orbital contribution to magnetic moments, application of magnetic moment data for characterization of 3d-metal complexes.

III. Thermodynamic and Kinetic Aspects of Metal Complexes (9)

A brief outline of thermodynamic stability of metal complexes and factors affecting the stability, substitution reactions of square planar complexes.

Electronic Spectra of Transition Metal Complexes Term Symbols for p² & d² systems, spectroscopic ground states for d¹-d¹⁰ electronic configurations. Types of electronic transitions, selection rules for d-d transitions, spectroscopic ground states, Orgel diagram for d¹-d⁵.

IV. Organometallic Compounds (8)

Definition, nomenclature and classification of organometallic compounds. EAN rule, Preparation, properties, and applications of alkyls aryls of lithium and aluminium, Bonding in metal-ethylenic complexes, Mechanism of homogeneous hydrogenation reactions.

Course Outcome: At the end of the course, students will be able to

1. Use valence bond theory and crystal field theory to predict the structure and magnetic behaviour of metal complexes.
2. Explain and understand spin-only and effective magnetic moments and relate it to the structure of complexes.
3. Understand and differentiate between kinetic and thermodynamic stability.
4. Gain insight into various aspects related to electronic transitions.
5. Understand the chemistry of organometallic compounds.

Text Books

1.B.N. Figgis, Introduction to Ligand Field, Wiley Eastern.

FOR BATCH 2023 AND ONWARDS
SARDAR BEANT SINGH STATE UNIVERSITY GURDASPUR

2. A.B.P. Lever, Inorganic Electronic Spectroscopy, Elsevier.
3. A. Earnshaw, Introduction to Magnetochemistry, Academic Press.
4. J.E. Huheey, Inorganic Chemistry Principles of Structure and Reactivity, Harper Inter- Science.
5. R.S. Drago, Physical Method in Chemistry, W.B. Saunders Company.
6. F.A. Cotton and G. Wilkinson, Advanced Inorganic Chemistry, Wiley Inter-science.

BSNM-23502
PHYSICAL CHEMISTRY-IV

L T P C
3 0 0 3

Course Objective

This program is intended to provide the students an in depth understanding of basic as well as advance concepts of electrochemistry. Detailed discussion of nuclear chemistry including various laws governing the nuclear processes and various factors affecting them will be discussed along with brief introduction to spectroscopy discussing rotational and vibrational spectroscopy and electronic spectrum.

I. Electrochemistry-I

(9)

Electrical transport-conduction in metals and in electrolyte solutions, specific conductance and equivalent conductance, measurement of equivalent conductance, variation of equivalent and specific conductance with dilution. Migration of ions and Kohlrausch law, Arrhenius theory of electrolyte dissociation and its limitations, weak and strong electrolytes, Ostwald's dilution law, its uses and limitations. Debye-Huckel-Onsager's equation for strong electrolytes (elementary treatment only). Transport number, definition and determination by Hittorf method and moving boundary method. Applications of conductivity measurements: determination of degree of dissociation, determination of K_a of acids, determination of solubility product of a sparingly soluble salt, conductometric titrations.

II. Electrochemistry – II

(9)

Types of reversible electrodes-gas metal ion, metal ion, metal insoluble salt-anion and redox electrodes. Electrode reactions. Nernst equation, derivation of cell E.M.F. and Single electrode potential, standard hydrogen electrode, reference electrodes, standard electrode potential, sign conventions, electrochemical series and its significance. Electrolytic and Galvanic cells-reversible and irreversible cells, conventional representation of electrochemi cells. EMF of a cell and its measurements. Computation of cell. EMF, Calculation of thermodynamic quantities of cell reactions (ΔG ΔH and K), polarization, over potential and hydrogen overvoltage. Concentration cells with and without transport, liquid junction potential, application of concentration cells, valency of ions, solubility product and activity coefficient, potentiometric titrations. Definition of pH and pKa, determination of pH using hydrogen, quinhydrone and glass electrodes, by potentiometric methods. Buffers-mechanism of buffer action, Henderson-Hassel equation, Hydrolysis of salts. Corrosion-types, theories and methods of combating it.

III. Nuclear Chemistry **(9)**

Introduction: Radioactivity, Nuclear Structure, Size of Nucleus, Mass Defects and Binding Energy, Nuclear Stability, Nuclear Forces, Nuclear Spin and Moments of Nuclei, Nuclear Models, Nuclear Decay Processes, The Laws of Radioactive Decay, Soddy-Fajans Group Displacement Law, Rate of Nuclear Decay and Half Life Time (Kinetics of Radioactive Decay), Induced Nuclear Reactions, Types of Nuclear Processes, High Energy Nuclear Reactions, Nuclear Reaction Cross-Section, Artificial radioactivity, Detection and Measurement of Radioactivity, Nuclear Fission, Nuclear Fusion, Applications of Radioactivity.

IV. Spectroscopy **(8)**

Introduction: Electromagnetic radiation, regions of the spectrum, basic features of different spectrometers, statement of the Born-Oppenheimer approximation, degrees of freedom.

Rotational Spectrum Diatomic molecules. Energy levels of a rigid rotor (semiclassical principles), selection rules, spectral intensity, distribution using population distribution (Maxwell-Boltzmann distribution) determination of bond length, qualitative description of non-rigid rotor, isotope effect.

Vibrational Spectrum Infrared spectrum: Energy levels of simple harmonic oscillator, selection rules, pure vibrational spectrum, intensity, determination of force constant and qualitative relation of force constant and bond energies, effect of anharmonic motion and isotope on the spectrum, idea of vibrational frequencies of different functional groups. Raman Spectrum: Concept of polarizability, pure rotational and pure vibrational Raman spectra of diatomic molecules, selection rules.

Electronic Spectrum Concept of potential energy curves for bonding and antibonding molecular orbitals, qualitative description of selection rules and Franck-Condon principle. Qualitative description of s, p, and n M.O., their energy levels and the respective transitions.

Course Outcome

1. To understand the interconversion of chemical and electrical energy and to link thermodynamics with electrochemistry.
2. To apply the concept of electrochemistry, spectroscopy to different chemical processes as well as in practicals.
3. The students will be able to pursue their carrier objectives in higher education, scientific research and teaching.

Text Books

1. Thermodynamics for Chemists, S. Glasstone.
2. R.S. Drago, "Physical Methods in Chemistry".
3. Principles of Physical Chemistry, S.H. Maron & C.F. Prutton.
4. Physical Chemistry, P.W. Atkins.
5. G.M. Barrow "Introduction to Molecular Spectroscopy".
6. C.N. Banwell "Fundamentals of Molecular Spectroscopy".
7. Concise Inorganic Chemistry by J.D. Lee, Oxford; Fifth edition).

BSNM-23503
ELEMENTS OF MODERN PHYSICS

L	T	P	C
4	0	0	4

Course Objective: By virtue of this course, the students will be able to understand some basic concepts and principles of quantum mechanics, atomic physics and nuclear physics.

I. Dual Nature of Waves and Particles (15)

Black body radiation, Planck's quantum, Planck's constant and light as a collection of photons; Photo Electric effect and Compton scattering. de-Broglie wavelength and matter waves; Davisson-Germer experiment, Wave-particle duality, Heisenberg uncertainty principle- Position-momentum and Energy-time uncertainty principle. Applications of uncertainty principle; Impossibility of a particle following a trajectory; Estimating minimum energy of a confined particle using uncertainty principle.

II. Quantum Mechanics (15)

Two slit interference experiment with photons, atoms & particles; linear superposition principle as a consequence; Matter waves and wave amplitude; Schrodinger wave equation; Momentum and Energy operators; Wavefunction, Physical interpretation of wavefunction, Probabilities, Normalization of wavefunction, Orthogonal wavefunction; One dimensional infinitely rigid box-energy eigenvalues and eigenfunctions.

III. Atomic structure (15)

Problems with Rutherford model- instability of atoms and observation of discrete atomic spectra, The Bohr Model of atom, Bohr's theory of Hydrogen atom, Energy level and spectra of Hydrogen atom, Limitations of Bohr's theory, Correspondence principle, Atomic excitation and ionization potentials, Production of X-rays and their spectra. Moseley's law, X-ray absorption.

Course Outcome: Upon successful completion of this course, students will be able to:

1. Understand and explain the differences between classical and quantum mechanics.
2. Discuss quantum mechanical problems dealing with the Heisenberg uncertainty principle
3. Solve Schrodinger equation for simple potentials.

Text Books

1. Concepts of Modern Physics, Arthur Beiser, 2009, McGraw-Hill
2. Modern Physics, J.R. Taylor, C.D. Zafiratos, M.A. Dubson, 2009, PHI Learning
3. Six Ideas that Shaped Physics: Particle Behave like Waves, Thomas A. Moore, 2003, McGraw Hill
4. Quantum Physics, Berkeley Physics, Vol.4. E.H. Wichman, 2008, Tata McGraw-Hill Co.

BSNM-23504
QUANTUM MECHANICS

L	T	P	C
4	0	0	4

Course Objective

The aim of this course is to make students acquire sufficient knowledge in Properties of matter waves, Wave equation, Schrodinger equation and applications, Operator formalism.

I. Time dependent and independent Schrodinger equation (10)

Time dependent Schrodinger equation, dynamical evolution of a quantum state; Interpretation of Wave Function, Probability and probability current densities in three dimensions; Conditions for Physical Acceptability of Wave Functions. Position, momentum & Energy operators; Expectation value, Commutator of position and momentum operators; Wave Function of a Free Particle. Time independent Schrodinger equation, Hamiltonian, stationary states and energy eigenvalues; General solution of the time dependent Schrodinger equation, wave packets, Fourier transforms and momentum space wave function; Position-momentum uncertainty principle.

II. Applications of Schrodinger Equation (10)

General discussion of bound states in an arbitrary potential- continuity of wave function, boundary condition and emergence of discrete energy levels; application to one-dimensional problem-square well potential; Quantum mechanics of simple harmonic oscillator-energy levels and energy eigenfunctions using Frobenius method.

III. Quantum theory of hydrogen-like atoms (10)

Time independent Schrodinger equation in spherical polar coordinates; separation of variables for the second order partial differential equation; angular momentum operator and quantum numbers; Radial wavefunctions from Frobenius method; Orbital angular momentum quantum numbers l and m ; s, p, d,... shells (idea only).

IV. Atoms in Electric and Magnetic Fields (10)

Electron Angular Momentum. Space Quantization. Electron Spin and Spin Angular Momentum. Larmor's Theorem. Spin Magnetic Moment. Stern-Gerlach Experiment. Zeeman Effect: Electron Magnetic Moment & Magnetic Energy, Gyromagnetic Ratio & Bohr Magneton. Atoms in External Magnetic Fields: Normal and Anomalous Zeeman Effect.

Course Outcome: On the successful completion of the course, students will be able to

1. Recognize basic terms in Quantum Mechanics
2. Understand the basic principles of quantum particles
3. Apply basics to construct and solve one particle equations
4. Ability to design and construct particle equations in the free and bound states as well as to analyze and interpret the results

5. To understand the fundamentals and concepts in operator formalism

Text Books

1. A Text book of Quantum Mechanics, P.M.Mathews & K.Venkatesan, 2nd Ed., 2010, McGraw Hill
2. Quantum Mechanics, Robert Eisberg and Robert Resnick, 2nd Edn., 2002, Wiley.
3. Quantum Mechanics, Leonard I. Schiff, 3rd Edn. 2010, Tata McGraw Hill.
4. Quantum Mechanics, G. Aruldas, 2nd Edn. 2002, PHI Learning of India.
5. Quantum Mechanics, Bruce Cameron Reed, 2008, Jones and Bartlett Learning.
6. Quantum Mechanics for Scientists and Engineers, D.A.B. Miller, 2008, Cambridge University Press
7. Quantum Mechanics, Eugen Merzbacher, 2004, John Wiley and Sons, Inc.
8. Introduction to Quantum Mechanics, David J. Griffith, 2nd Ed. 2005, Pearson Education
9. Quantum Mechanics, Walter Greiner, 2nd Edn., 2001, Springer

BSNM-23505
THEORY OF PROBABILITY

L	T	P	C
3	0	0	3

Course Objective

The objective of the course is to prepare students for big data analysis by introducing basic concepts of probability theory along with their applications.

I. (9)

Random experiment, sample space, event, algebra of events, Probability definition, addition law of probability, multiplication law of probability, conditional probability and independence, Bayes' Theorem.

II. (9)

Random variables, distribution function, properties of distribution function, discrete random variable, probability mass function, discrete distribution function, continuous random variable, probability density function. Continuous distribution function.

III. (8)

Mathematical expectation, expectation of a random variable, Discrete probability distributions: binomial, Poisson, negative binomial distribution.

IV. (9)

Continuous probability distributions: uniform distribution, normal distribution, normal distribution as a limiting case of binomial distribution, Gamma distribution, Beta distribution.

Course Outcome: At the end of the course, the students will be able to

1. Understand the mathematical definition of probability, conditional probability and its applications.
2. Understand the theoretical concepts like random variable, probability distribution, generating functions and their usage

Text Books

1. S. Ross, A First Course in Probability, Pearson.
2. S.C. Gupta, V. K. Kapoor, Fundamentals of Mathematical Statistics, Sultan Chand & Sons, Delhi

BSNM-23506
NUMERICAL ANALYSIS

L	T	P	C
3	0	0	3

Course Objective

This course is aimed at providing the students with knowledge about the numerical solutions to various mathematical expressions which are not easily solvable by conventional techniques. These techniques are very useful for the students for experimental data analysis, integration and differentiation of involved functions, solutions of certain implicit equations.

I. Linear System of Equations (8)

Gauss elimination method, Gauss Jordan method, LU decomposition method. Iterative Methods: Jacobi, Gauss-Seidel, Relaxation Methods; Eigenvalue Problem: Power Method.

II. Interpolation (9)

Interpolation with Unevenly Spaced Points: Lagrange Interpolation, Newton's Divided Difference Interpolation; Interpolation with Evenly Spaced Points: Newton's Forward Difference Interpolation Formula, Newton's Backward Difference Interpolation Formula, Spline interpolation.

III. Numerical Differentiation and Integration (9)

Numerical differentiation: Newton's Forward Difference Formula, Newton's Backward Difference Formula, Newton's Divided Difference Formula; Numerical Integration: Trapezoidal rule, Simpson's 1/3-rule and Simpson's 3/8 rule.

IV. Numerical solution of ordinary differential equations (ODEs) (9)

Initial Value Problems of ODEs: Taylor series method, Euler's methods, Runge-Kutta methods and linear multi-step methods (Adams-Bashforth & Adams-Moulton).

Course Outcome: At the end of the this course, students will be able to

1. Learn the application of Numerical methods and its application
2. Understand the concept of errors and their significance in numerical methods
3. Learn different numerical methods used for solution of linear and non linear equations
4. Learn how different types of problems can be solved by using different methods

Text Books

1. Richard L. Burden and J. Douglas Faires, Numerical Analysis, 9th Edition, Cengage Learning
2. M. K. Jain, S. R. K. Iyengar and R. K. Jain, Numerical Methods for Scientific and Engineering Computation, 6th Edition, New Age International Publisher

BSNM-23507
ENGLISH-V

L T P C
2 0 0 2

Course Objective

Educating students to use the rules of English grammar. Training them to produce different grammatical structures such statements and questions correctly.

1. Literature

(6)

The Poetic Palette (Orient Black Swan, Second Edition, 2016)

The following poems from this anthology are prescribed:

- a. The Charge of the Light Brigade: Alfred Tennyson
- b. He Wishes for the Cloths of Heaven: W. B. Yeats
- c. True ease in writing comes from art, not chance: Alexander Pope
- d. Goodbye party for Miss Pushpa T. S.: Nissim Ezekiel

2. Vocabulary

(4)

Various processes of Word formation; Standard Abbreviations & Acronyms; Internet Texting Abbreviations & Acronyms

3. Literature

(6)

Prose Parables (Orient Black Swan, 2013)

The following stories from the above volume are prescribed:

- a. The Voice of God: Prem Chand
- b. The Face on the Wall: E.V. Lucas
- c. The Gold Frame: R. K. Laxman
- d. My Brother, My Brother: Norah Burke

4. Grammar

(9)

Use of Idioms/Phrases in sentences; Understanding Sentences Structures & practice on Transformation of sentences.

Reading & Writing Skills:

Close Reading; Comprehension; Translation (from Hindi/Punjabi to English and vice-versa)

Business Correspondence- Business letters; Letter to the Editor; Business Emails; Drafting Notices & Memos

Interactive practice sessions on Oral Communication

Self-Introduction, Group Discussion and Role Play

Common Everyday Situations: Conversations and Dialogues

Course Outcome

1. Understand grammatical concepts of verb phrases.

2. Infer the main theme and other ideas in the text studies

Text Books

1. Oxford Practice Grammar by John Eastwood (Ed. 2014)
2. Business English, Pearson, 2008.
3. Language, Literature and Creativity, Orient Black swan, 2013.
4. Remedial English Grammar. F.T. Wood. Macmillan.2007.
5. On Writing Well. William Zinsser. Harper Resource Book. 2001
6. Study Writing. Liz Hamp-Lyons and Ben Heasley. Cambridge University Press. 2006.
7. Exercises in Spoken English. Parts. I-III. CIEFL, Hyderabad. Oxford University Press.

BSNM-23508
PUNJABI-V

L T P C
2 0 0 2

ਪਾਠ-ਕ੍ਰਮ:

ਯੂਨਿਟ-1 (ਸਾਹਿਤ)

1. ਡਾ. ਗੰਡਾ ਸਿੰਘ - ਪ੍ਰੋ. ਪ੍ਰੀਤਮ ਸਿੰਘ
2. ਨਾਨਕ ਸਿੰਘ - ਬਲਵੰਤ ਗਾਰਗੀ
3. ਬਾਬਾ, ਬੋਹੜ ਨਹੀਂ - ਭਗਵੰਤ ਸਿੰਘ
4. ਨਿੱਕੀ ਕਹਾਣੀ ਦਾ ਬਾਦਸ਼ਾਹ-ਅਜੀਤ ਕੌਰ
5. ਬਾਤਾਂ ਮੋਹਨ ਸਿੰਘ ਕੀਆਂ- ਕੁਲਬੀਰ ਸਿੰਘ ਕਾਂਗ
6. ਗੁਲਾਬੀ ਕਾਗਜ਼ ਉੱਤੇ ਲਿਖੀ ਕਵਿਤਾ:ਸੰਤੋਖ ਧੀਰ-ਗੁਰਬਚਨ ਸਿੰਘ ਭੁੱਲਰ
7. ਸੁਤਿੰਦਰ ਸਿੰਘ ਨੂਰ: ਸਾਹਿਤ ਦਾ ਜਥੇਦਾਰ-ਗੁਰਬਚਨ
8. ਮਿਲਖਾ ਸਿੰਘ-ਸਰਵਣ ਸਿੰਘ

ਯੂਨਿਟ-2 (ਭਾਸ਼ਾ)

ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਵਿਚ ਆਏ ਪਰਵਿਰਤਨ
ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਦੀ ਵਿਗਿਆਨ ਦੀ ਸਿਖਿਆ ਵਿਚ ਭੂਮਿਕਾ

ਯੂਨਿਟ-3 (ਵਿਆਕਰਣ)

ਪੰਜਾਬੀ ਵਿਆਕਰਣਕ ਇਕਾਈਆਂ: ਸਵਾਧੀਨ ਉਪਵਾਕ ਤੇ ਪਰਾਧੀਨ ਉਪਵਾਕ।

ਯੂਨਿਟ-4 (ਲੇਖਣੀ-ਕਲਾ)

ਸਨੇਹੀਆਂ ਨੂੰ ਚਿੱਠੀ-ਪੱਤਰ
ਪੋਸਟ ਕਾਰਡ ਲਿਖਣ ਦੀ ਵਿਧੀ ਤੇ ਨਮੂਨਾ

FOR BATCH 2023 AND ONWARDS
SARDAR BEANT SINGH STATE UNIVERSITY GURDASPUR

ਸਹਾਇਕ ਪੁਸਤਕਾਂ:

ਸਾਹਿਤ ਦੇ ਰੰਗ (ਸੰਪ. ਡਾ.ਮਹਿਲ ਸਿੰਘ), ਰਵੀ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ ਅੰਮ੍ਰਿਤਸਰ।

ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਦਾ ਵਿਆਕਰਨ ਜੋਗਿੰਦਰ ਸਿੰਘ ਪੁਆਰ, ਬਲਦੇਵ ਸਿੰਘ ਚੀਮਾ, ਸੁਖਵਿੰਦਰ ਸਿੰਘ ਸੰਘਾ, ਵੇਦ ਅਗਨੀਹੋਤਰੀ), ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਅਕਾਦਮੀ, ਜਲੰਧਰ, ਐਡੀਸ਼ਨ 2009.

BSNM-23509

Drug Abuse-I (Problem, and Management)

L T P C
2 0 0 0

Course Objective

This course will cover some basic information about the nature and effects of drugs of abuse. The course will also address the social and personal dynamics involved in the phenomena of drug abuse. Finally, the various measures for prevention currently being employed to combat drug abuse will be explored.

I. Meaning of Drug Abuse

(6)

Concept and Overview, Historical Perspective of Drug Abuse, Drug Dependence, Drug Addiction, Physical and Psychological Dependence: Drug Tolerance and withdrawal symptoms.

II. Types of Abused Drugs and their Effects

(7)

- 1) Stimulants: Amphetamines – Benzedrine, Dexedrine, Cocaine.
- 2) Depressants: Alcohol Barbiturates: Nembutal, Seconal, Phenobarbital and Rohypnol.
- 3) Narcotics: Heroin, Morphine, Oxycodone.
- 4) Hallucinogens: Cannabis, Marijuana, Hashish, Hash Oil, MDMA, LSD.
- 5) Steroids.

III. Nature and Extent of the Problem

(6)

Magnitude or prevalence of the menace of Drug Abuse in India and Punjab, Vulnerable groups by age, gender and economic status, Signs and Symptoms of Drug Abuse: Physical, Academic, Behavioural and Psychological Indicators.

IV. Management of Drug Abuse

(6)

Medical Management: Medication for treatment and to reduce withdrawal effects. Psychiatric Management: Counselling, Behavioural and Cognitive therapy. Social Management: Family, Group therapy and Environmental Intervention.

Course Outcome: Students will be able to

1. Identify some of the current trends in subject matter.
2. Recognize the impact drugs have on individuals and society as a whole.
3. Describe the action of drugs and the role the drug has had in societies of the past and in the current world.
4. Classify drugs into the drug category to which they belong, as well as identify the schedule the drug is assigned in this country.

5. Evaluate the pros and cons of substance use, society's view of different substances, and the pros and cons of current drug law.

Text Books

1. Ahuja, Ram (2003), Social Problems in India, Rawat Publication, Jaipur.
2. Extent, Pattern and Trend of Drug Use in India, Ministry of Social Justice and Empowerment, Government of India, 2004.
3. Inciardi, J.A. 1981. The Drug Crime Connection. Beverly Hills: Sage Publications.
4. Kapoor. T. (1985) Drug epidemic among Indian Youth, New Delhi: Mittal Pub. 15
5. Modi, Ishwar and Modi, Shalini (1997) Drugs: Addiction and Prevention, Jaipur: Rawat Publication.
6. National Household Survey of Alcohol and Drug abuse. (2003) New Delhi, Clinical Epidemiological Unit, All India Institute of Medical Sciences, 2004.
7. Sain, Bhim 1991, Drug Addiction Alcoholism, Smoking obscenity New Delhi: Mittal Publications.
8. Singh, Chandra Paul 2000. Alcohol and Dependence among Industrial Workers: Delhi: Shipra.
9. Sussman, S and Ames, S.L. (2008). Drug Abuse: Concepts, Prevention and Cessation, Cambridge University Press.

BSNM-23510
CHEMISTRY LAB-V

L	T	P	C
0	0	4	2

1. Synthesis and Analysis

- (a) Preparation of Sodium trioxalatoferrate(III)
- (b) Preparation of Ni-DMG Complex
- (c) Preparation of Copper tetrammine complex
- (d) Preparation of cis-bisoxalatodiaquachromate(III)ion

2. Physical Chemistry

(a) Conductometric Titrations

(i) Determine the end point of the following titrations by the conductometric methods.

Strong acid-Strong base

Strong acid-Weak base

Weak acid-Strong base

Weak acid-Weak base

(ii) Determine the composition of a mixture of acetic acid and the hydrochloric acid by conductometric titration.

3. Molecular Weight Determination of acetanilide, naphthalene, using camphor as solvent (Rast's methods).

4. To determine the molecular weight of a polymer by viscosity measurements.

5. Adsorption: To study the adsorption of acetic acid oxalic/acid from aqueous solutions by charcoal.

6. Phase Equilibria: To determine the distribution coefficient of iodine between CCl₄ and water.

7. Refractometry: (i) Determination of refractive index of a liquid by Abbe refractometer, and hence the specific and molar refraction.

(ii) To determine the composition of unknown mixture of two liquids by refractive index measurements.

Text Books

- 1. Practical Inorganic Chemistry by J.R. Barrante G. Marr and B.W. Rockett
- 2. Vogel's Inorganic Quantitative Analysis
- 3. Advanced Practical Physical Chemistry by J.B. Jadav

BSNM-23511
PHYSICS LAB -V

L	T	P	C
0	0	4	2

Students are required to perform at least 06 experiments from the following:

1. Measurement of Planck's constant using black body radiation and photo-detector.
2. Photo-electric effect: photo current versus intensity and wavelength of light; maximum energy of photo-electrons versus frequency of light.
3. To determine work function of material of filament of directly heated vacuum diode.
4. To determine the Planck's constant using LEDs of at least 4 different colours.
5. To determine the wavelength of H-alpha emission line of Hydrogen atom.
6. To determine the ionization potential of mercury.
7. To determine the absorption lines in the rotational spectrum of Iodine vapour.
8. To determine the value of e/m by (a) Magnetic focusing or (b) Bar magnet.
9. To setup the Millikan oil drop apparatus and determine the charge of an electron.
10. To show the tunneling effect in tunnel diode using I-V characteristics.
11. To determine the wavelength of laser source using diffraction of single slit.
12. To determine the wavelength of laser source using diffraction of double slits.
13. To determine (1) wavelength and (2) angular spread of He-Ne laser using plane diffraction grating.
14. Study of Electron spin resonance- determine magnetic field as a function of the resonance frequency.
15. Study of Zeeman effect: with external magnetic field; Hyperfine splitting.
16. To study the quantum tunnelling effect with solid state device, e.g. tunnelling current in backward diode or tunnel diode.

Text Books

1. Advanced Practical Physics for students, B. L. Flint and H.T. Worsnop, 1971, Asia Publishing House.
2. Advanced level Physics Practicals, Michael Nelson and Jon M. Ogborn, 4th Edition, reprinted 1985, Heinemann Educational Publishers
3. Engineering Practical Physics, S.Panigrahi & B. Mallick, 2015, Cengage Learning India Pvt. Ltd.
4. Practical Physics, G.L. Squires, 2015, 4th Edition, Cambridge University Press.
5. A Text Book of Practical Physics, I.Prakash & Ramakrishna, 11th Edn, 201, Kitab Mahal.
6. B Sc. Practical Physics, C. L. Arora, S. Chand & Co.
7. A Text Book of Practical Physics, Indu Prakash and Ramakrishna, 11th Edition, 2011, Kitab Mahal, New Delhi.

BSNM- 23601
ORGANIC CHEMISTRY-IV

L T P C
4 0 0 4

Course Objective

The objective of this course is to study the various aspects of transition metal complexes and organometallic compounds in detail. This will help the student to apply the knowledge in the advanced field of inorganic chemistry.

I. Spectroscopy

(10)

Nuclear Magnetic Resonance (NMR) spectroscopy. Proton Magnetic Resonance (¹H NMR) spectroscopy, nuclear shielding and deshielding, chemical shift and molecular structure, spin-spin splitting and coupling constants, areas of signals, interpretation of PMR spectra of simple organic molecules such as ethyl bromide, ethanol, acetaldehyde, 1,1,2-tribromoethane, ethyl acetate, toluene and acetophenone.

Electromagnetic Spectrum: Absorption Spectroscopy Ultraviolet (U.V.) absorption spectroscopy introduction- (Beer-Lambert law), molar absorptivity, analysis of UV spectra, types of electronic transitions effect of conjugation. Concept of chromophores and auxochrome, Bathochrome, hypsochrome, hyperchrome, hypochromic shifts-UV spectra of conjugated compounds, Infrared (IR) Absorption spectroscopy-introduction, Hooke's law, Selection rules, intensity and IR bands, measurement of IR spectrum time characteristic absorption of various fundamental band interpretation of IR spectra of simple organic compounds.

Problems pertaining to the structure elucidation of simple organic compounds using UV, IR and PMR spectroscopic techniques. Synthetic Polymers Addition or chain-growth polymerization. Free radical vinyl polymerization, ionic vinyl polymerization, Ziegler-Natta polymerization and vinyl polymers. Condensation or step growth polymerization. Polyesters, polyamides, phenol formaldehyde resins, urea formaldehyde resins, epoxy resins and polyurethanes. Natural and synthetic rubbers.

II. Organosulphur Compounds

(10)

Nomenclature, structural features, Methods of formation and chemical reactions of thiols, thioethers, sulphonic acids, sulphonamides and sulphaguanidine. Organic Synthesis via Enolates Acidity of α -hydrogens, alkylation of diethyl malonate and ethyl acetoacetate. Synthesis of ethyl acetoacetate: the Claisen condensation. Keto-enol tautomerism of ethyl acetoacetate. Alkylation of 1,3-dithianes. Alkylation and acylation of enamines.

III. Carbohydrates

(10)

Classification and nomenclature. Monosaccharides, mechanism of osazone formation, interconversion of glucose and fructose, chain lengthening and chain shortening of aldoses.

Configuration of monosaccharides. Erythro and threo diastereomers. Conversion of glucose into mannose. Formation of glycosides, ethers and esters. Determination of ring size of monosaccharides. Cyclic structure of D(+)-glucose. Mechanism of mutarotation. Structures of ribose and deoxyribose. An introduction to disaccharides (maltose, sucrose and lactose) and polysaccharides (starch and cellulose) without involving structure determination.

IV. Amino Acids, Peptides, Proteins and Nucleic Acids **(10)**

Classification, structure and stereochemistry of amino acids. Acid-base behaviour, isoelectric point and electrophoresis. Preparation and reactions of α -amino acids. Structure and nomenclature of peptides and proteins. Classification of proteins. Peptide structure determination, end group analysis, selective hydrolysis of peptides. Classical peptide synthesis, solid-phase peptide synthesis. Structures of peptides and proteins. Levels of protein structure.

Protein denaturation/renaturation. Nucleic acids: Introduction. Constituents of nucleic acids. Ribonucleosides and ribonucleotides. The double helical structure of DNA.

Course Outcome: At the end of the course, students will be able to

1. Use valence bond theory and crystal field theory to predict the structure and magnetic behaviour of metal complexes.
2. Explain and understand spin-only and effective magnetic moments and relate it to the structure of complexes.
3. Understand and differentiate between kinetic and thermodynamic stability.
4. Gain insight into various aspects related to electronic transitions.
5. Understand the chemistry of organometallic compounds.

Text Books

1. Organic Chemistry. F.A. Carey, McGraw Hill, Inc. 8th edition.
2. Organic Chemistry, Morrison and Boyd, Prentice Hall
3. R.M. Silverstein, G.C. Bassler, T.C. Morrill, "Spectrometric Identification of Organic Compounds.
4. W. Kemp, "Organic Spectroscopy".
5. D.H. Williams, I. Fleming, "Spectroscopic Methods in Organic Chemistry".
6. J.R. Dyer, "Application of Absorption Spectroscopy of Organic Compounds".
7. D. H. Williams, I. Fleming, "Spectroscopic Problems in Organic Chemistry" 1967.
8. R.C. Banks, E.R. Matjeka, G. Mercer, "Introductory Problems in Spectroscopy" 1980.

BSNM- 23602
PHYSICAL CHEMISTRY-IV

L	T	P	C
3	0	0	3

Course Objective

This program is intended to provide the students an in depth understanding of basic as well as advance concepts of electrochemistry. Detailed discussion of nuclear chemistry including various laws governing the nuclear processes and various factors affecting them will be discussed along with brief introduction to spectroscopy discussing rotational and vibrational spectroscopy and electronic spectrum

I. Quantum Mechanics-I

(12)

Black-body radiation, Planck's radiation law, Photoelectric effect, heat capacity of solids, Bohr's model of hydrogen atom (no derivation) and its defects, Compton effect. de Broglie hypothesis, Heisenberg's uncertainty principle, Sinusoidal wave equation, Hamiltonian operator, Schrodinger wave equation and its importance, physical interpretation of the wave function, postulates of quantum mechanics, particle in a one dimensional box, quantization of energy levels, extension to two and three dimensional boxes, degeneracy.

II. Quantum Mechanics-II

(12)

Simple harmonic oscillator model of vibrational motion, setting up Schrodinger equation and discussion of solution and wave functions. Rigid rotator model of rotation of diatomic molecules transformation to spherical polar coordinates spherical harmonics and their discussion. Qualitative investigation H-atom, setting up Schrodinger equation, radial and angular part, radial distribution functions of 1s, 2s, 2p, 3s, 3p and 3d.

III. Photochemistry

(11)

Interaction of radiation with matter, difference between thermal and photochemical processes. Laws of photochemistry: Grothus–Drapper law, Stark–Einstein law, Jablonski diagram depicting various processes occurring in the excited state, qualitative description of fluorescence, phosphorescence, non–radiative processes (internal conversion, intersystem crossing), quantum yield, photosensitized reactions–energy transfer processes (simple examples).

Course Outcome

1. To understand the interconversion of chemical and electrical energy and to link thermodynamics with electrochemistry.

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2. To apply the concept of electrochemistry, spectroscopy to different chemical processes as well as in practicals.
3. The students will be able to pursue their carrier objectives in higher education, scientific research and teaching.

Text Books

1. Physical Chemistry, A Molecular Approach by D.A. Mcguarrie and J.D. Simon.
2. Quantum Chemistry, Ira N. Levine.
3. Quantum Chemistry, H. Eyring J. Walter and G.E. Kimball.
4. Molecular Quantum Mechanics, P.W. Atkins.
5. R.S.Drago, "Physical Methods in Chemistry".

BSNM- 23603
SOLID STATE PHYSICS

L	T	P	C
4	0	0	4

Course Objective: The course gives an introduction to solid state physics, and will enable the student to employ classical and quantum mechanical theories needed to understand the physical properties of solids.

1. Crystal Structure (12)

Basis and crystal structure, Unit Cell, Lattice vectors, Two dimensional lattice, Three dimensional lattices, Symmetry operations, Miller indices, Cubic structures, Hexagonal close packed structure. Interplanar spacing, Diffraction of X-rays, Bragg's law of diffraction. Reciprocal lattice, Reciprocal lattice to SC, BCC and FCC lattice, Brillouin zone, Atomic form factor, geometrical structure factor.

2. Lattice Vibrations (10)

Lattice vibrations of 1D monoatomic lattice and diatomic lattice. Phonons and their momentum. Inelastic scattering of photons by phonons. Specific heat, classical theory of lattice heat capacity, Einstein's theory of lattice heat capacity, Debye model of lattice heat capacity, Debye approximation, Limitations of Debye model.

3. Free Electron Theory (10)

Drude-Lorentz theory, Electrical conductivity and Ohm's Law, Sommerfeld model, Fermi-Dirac distribution function, Effect of temperature on f-d distribution, Thermal conductivity of metals. Wiedemann -Frenz law, Hall effect.

4. Band Theory (10)

Origin and magnitude of energy band gap, Density of states, Bloch theorem, Kronig-Penney model of an infinite one dimensional crystal, Classification of insulators, semiconductors and metals. The tight-binding approximation in evaluating the energy levels for an electron in a solid. Direct and indirect energy band semiconductors.

Course Outcome: At the end of this course:

1. Students will be able to classify various types of structures and correlate various properties of solid materials.
2. Students will be able to know the concept of phonons, heat capacity etc.

3. Students will be able to classify metals, semiconductors and insulators on the basis of band theory.

Text Books

1. Introduction to Solid State Physics, Charles Kittel, 8th Ed., 2004, Wiley India Pvt. Ltd.
2. Elements of Solid State Physics, J.P. Srivastava, 2nd Ed., 2006, Prentice-Hall of India
3. Introduction to Solids, Leonid V. Azaroff, 2004, Tata Mc-Graw Hill
4. Solid State Physics, N.W. Ashcroft and N.D. Mermin, 1976, Cengage Learning
5. Solid-state Physics, H. Ibach and H. Luth, 2009, Springer
6. Elementary Solid State Physics, 1/e M. Ali Omar, 1999, Pearson India
7. Solid State Physics, M.A. Wahab, 2011, Narosa Publications

BSNM- 23604
NUCLEAR AND PARTICLE PHYSICS

L	T	P	C
4	0	0	4

Course Objective

The aim of this course is to provide a coherent and concise coverage of traditional nuclear physics.

I. Structure and Properties of the Nucleus (10)

Structure of the nucleus: Discovery of the nucleus, composition, basic properties; charge, mass, size, spin, magnetic moment, electric quadrupole moment, binding energy, binding energy per nucleon and its observed variation with mass number of the nucleus, coulomb energy, volume energy, surface energy, other corrections, explanation of the binding energy curve, liquid drop model of the nucleus, evidence for nuclear shell structure, nuclear magic numbers, basic assumption of shell model, concept of mean field, residual interaction, nuclear force.

II. Radioactive decays (10)

Alpha decay: basics of α -decay processes, theory of alpha emission, Gamow factor, Geiger Nuttall law, α -decay spectroscopy. (b) β -decay: energy kinematics for β -decay, positron emission, electron capture, neutrino hypothesis. (c) Gamma decay: Gamma rays emission & kinematics, internal conversion. Reactions: Types of Reactions, Conservation Laws, kinematics of reactions, Nuclear Q-value, reaction rate, reaction cross section, Concept of compound and direct reaction, resonance reaction, Coulomb scattering (Rutherford scattering).

III. Interaction of Radiation with Matter (10)

Energy loss of particles in passage through matter, stopping power of matter for charged particles, energy range relationship and straggling. Interaction of gamma radiation with matter: photoelectric effect, Compton effect and pair production. Thomson scattering and Rayleigh scattering. Detectors and Accelerators: Gas detectors: estimation of electric field, mobility of particle, for ionization chamber and GM Counter. Basic principle of Scintillation Detectors and construction of photo-multiplier tube (PMT). Semiconductor Detectors (Si and Ge) for charge particle and photon detection (concept of charge carrier and mobility), neutron detector, Need for accelerators.

IV. Cosmic Rays and Elementary Particles (10)

Discovery of cosmic rays: hard and soft components, discovery of elementary particle, muon, pion, heavy mesons and hyperons, mass and life time determination for muon and pion. Primary Cosmic Rays: Particle interactions; basic features, types of particles and its families. Symmetries and Conservation Laws: energy and momentum, angular momentum, parity, baryon number, Lepton number, Isospin, Strangeness and charm, concept of quark model, color quantum number and gluons.

Course Outcome: On the successful completion of the course, students will be able to

1. Acquire knowledge of the fundamental physics underpinning nuclear physics

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2. Apply general considerations of quantum physics to atomic and nuclear system
3. Analyze production and decay reactions for fundamental particles
4. Describe radioactivity and related phenomena
5. Explain the various interactions of nuclear radiation with matter

Text Books

1. R.D. Evans: Atomic Nucleus, Krieger Publishing Co. 2003
2. K.S. Krane: Introductory Nuclear Physics, Wiley 2008.
3. P. Mermier and E. Sheldon: Physics of Nuclei and particles, Academic Press, 2013.

BSNM-23605
MODERN ALGEBRA

L	T	P	C
3	0	0	3

Course Objective

A major objective is to introduce students to the language and precision of modern algebra. This means that the course will be proof-based, in the sense that students will be expected to understand, construct, and write proofs.

I. (9)

Groups, properties of group elements, subgroups, cyclic groups, cosets of a subgroup, Lagrange's theorem, normal subgroups and Quotient groups.

II. (9)

Homomorphism, Isomorphism theorems, conjugate elements, class equation, permutation groups, alternating groups, simplicity of A_n , $n \geq 5$ (without proof).

III. (9)

Rings, subring, characterization of a subring, integral domains, ideals, characteristic of a ring, Quotient rings.

IV. (8)

Prime and maximal Ideals, homomorphism, Isomorphism theorems, Polynomial rings.

Course Outcome: Upon successful completion of this course, students will be able to:

1. Effectively write abstract mathematical proofs in a clear and logical manner.
2. Locate and use theorems to solve problems in number theory and theory of polynomials over a field.
3. Demonstrate ability to think critically by interpreting theorems and relating results to problems in other mathematical disciplines.

Text Books

1. L. Gilbert, J. Gilbert, Elements of Modern Algebra, Cengage, 2015.
2. M. Artin, Algebra, Pearson, 2010

BSNM- 23606
ENGLISH - VI

L T P C
2 0 0 2

Course Objective

To train students to produce grammatically correct sentences. Train them to use verb tenses with appropriate time reference.

I . Literature

(6)

The study of the whole text of the play, All My Sons by Arthur Miller for vocabulary enrichment, learning sentence/speech construction and understanding dialogues/conversations.

II. Grammar and Vocabulary

(6)

Scientific/Technical Vocabulary; One word Substitution; Tenses; Active/Passive Voice; Narration; Common Errors .

III. Reading & Writing Skills

(6)

Summary & Paraphrasing, Analysis and Interpretation; Formal Report writing; Formal Presentations-Practice on preparing Formal Presentations; Power Point Presentations.

IV. Interactive practice sessions on Oral Communication

(7)

- Communication at Workplace
- Preparation for Interviews; Mock interviews
- Delivering Formal Presentations/Power Point Presentations/Oral Presentations.

Course Outcome

1. Students will heighten their their awareness of correct usage of English grammar in writing and speaking.
2. Students will attain and enhance competence in the four modes of literacy: writing, speaking, reading and listening.
3. Students will enlarge their vocabulary by keeping a vocabulary journal.

Text Books

1. Oxford Practice Grammar by John Eastwood (Ed. 2014)
2. Business English, Pearson, 2008.
3. Language, Literature and Creativity, Orient Black swan, 2013.
4. Remedial English Grammar. F.T. Wood. Macmillan. 2007.
5. On Writing Well. William Zinsser. Harper Resource Book. 2001
6. Study Writing. Liz Hamp-Lyons and Ben Heasley. Cambridge University Press. 2006
7. Exercises in Spoken English. Parts. I-III. CIEFL, Hyderabad. Oxford University Press.

BSNM- 23607
ENVIRONMENTAL SCIENCE

L T P C
2 0 0 2

Course Objective

To study about environment and ecosystems. To study about different types of natural resources. Knowledge and concept of biodiversity and its conservation. Basic knowledge and concept of causes, effect and control of different type of environmental pollution.

I. Introduction

(6)

Definition and scope and importance of multidisciplinary nature of environment. Need for public awareness. Ecosystems: Concept of Ecosystem, Structure, interrelationship, producers, consumers and decomposers, ecological pyramids-biodiversity and importance. Hot Spots of biodiversity, Natural Resources: Natural Resources and associated problems, use and over exploitation, case studies of forest resources and water resources.

II. Environmental Pollution

(6)

Definition, Causes, effects and control measures of air pollution, Water pollution, Soil pollution, Marine pollution, Noise pollution, Thermal pollution, Nuclear hazards. Solid waste Management: Causes, effects and control measure of urban and industrial wastes. Role of an individual in prevention of pollution. Pollution case studies. Disaster Management: Floods, earthquake, cyclone and landslides.

III. Social Issues: and the Environment

(6)

From Unsustainable to Sustainable development, Urban problems related to energy, Water conservation, rain water harvesting, watershed management. Resettlement and rehabilitation of people; its problems and concerns. Case studies. Environmental ethics: Issues and possible solutions. Climate change, global warming, acid rain, ozone layer depletion, nuclear accidents and holocaust. Case studies. Wasteland reclamation. Consumerism and waste products. Environment Protection Act. Air (Prevention and Control of Pollution) Act. Water (Prevention and control of pollution) Act. Wildlife Protection Act, Forest Conservation Act, Issues involved in enforcement of environmental legislation Public awareness.

IV. Human Population and the Environment

(6)

Population growth, variation among nations. Population explosion – Family Welfare Programme. Environment and human health, Human Rights, Value Education, HIV/AIDS. Women and child Welfare. Role of Information Technology in Environment and human health. Case studies

Field Work: Visit to a local area to document environmental assets river/ forest/ grassland/ hill/ mountain, Visit to a local polluted site-Urban/ Rural/ Industrial/ Agricultural, Study of common plants, insects, birds, Study of simple ecosystems-pond, river, hill slopes, etc. (Field work Equal to 5 lectures)

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Course Outcome: After completion of the course, a student will be able to

1. Gain knowledge about environment and ecosystem.
2. Students will learn about natural resource, its importance and environmental impacts of human activities on natural resource.
3. Aware students about problems of environmental pollution, its impact on human and ecosystem and control measures.
4. Students will learn about increase in population growth and its impact on environment

Text Books

1. Agarwal, K.C. 2001 Environmental Biology, Nidi Publ. Ltd. Bikaner.
2. Bharucha Erach, The Biodiversity of India, Mapin Publishing Pvt. Ltd., Ahmedabad – 3 013, India, Email:mapin@icenet.net (R)
3. Brunner R.C., 1989, Hazardous Waste Incineration, McGraw Hill Inc. 480p
4. Clark R.S., Marine Pollution, Clarendon Press Oxford (TB)
5. Cunningham, W.P. Cooper, T.H. Gorhani, E & Hepworth, M.T. 2001, Environmental Encyclopedia, Jaico Publ. House, Mumabai, 1196p
6. De A.K., Environmental Chemistry, Wiley Eastern Ltd.
7. Down to Earth, Centre for Science and Environment (R)

BSNM -23608
Punjabi-VI

L T P C
2 0 0 2

ਪਾਠ-ਕ੍ਰਮ:

ਯੂਨਿਟ-1 (ਸਾਹਿਤ)

1. ਕਿਰਤ - ਪ੍ਰੋ. ਪੂਰਨ ਸਿੰਘ
2. ਗੰਗਾ ਦੀਨ- ਪ੍ਰਿੰ. ਤੇਜਾ ਸਿੰਘ
3. ਮਾਂ-ਗੁਰਬਖਸ਼ ਸਿੰਘ ਪ੍ਰੀਤਲੜੀ
4. ਲਾਲ ਬਾਦਸ਼ਾਹ- ਹਰਿੰਦਰ ਸਿੰਘ ਰੂਪ
5. ਜਿਹੜੇ ਬੂਰੀਆਂ ਮੱਝੀਆਂ ਚੁੰਘਦੇ ਸੀ- ਸੂਬਾ ਸਿੰਘ
6. ਹਾਰ ਸਿੰਗਾਰ- ਗੁਲਜ਼ਾਰ ਸਿੰਘ ਸੰਧੂ
7. ਡੂੰਘੀਆਂ ਸਿਖਰਾਂ-ਨਰਿੰਦਰ ਸਿੰਘ ਕਪੂਰ
8. ਭਾਈ ਮਰਦਾਨਾ ਜੀ- ਹਰਪਾਲ ਸਿੰਘ ਪੰਨੂ

ਯੂਨਿਟ-2 (ਭਾਸ਼ਾ)

ਬਾਜ਼ਾਰ ਵਿਚ ਵਰਤੀ ਜਾਣ ਵਾਲੀ ਸ਼ਬਦਾਵਲੀ
ਵਪਾਰ ਵਿਚ ਵਰਤੀ ਜਾਣ ਵਾਲੀ ਸ਼ਬਦਾਵਲੀ

ਯੂਨਿਟ-3 (ਵਿਆਕਰਣ)

ਪੰਜਾਬੀ ਵਿਆਕਰਣਕ ਇਕਾਈਆਂ: ਨਾਂਵ ਵਾਕੰਸ਼ ਤੇ ਕਿਰਿਆ ਵਾਕੰਸ਼।

ਯੂਨਿਟ-4 (ਲੇਖਣੀ-ਕਲਾ)

ਅਖਬਾਰੀ ਲੇਖ
ਈ-ਮੇਲ ਲਿਖਣ ਦੀ ਵਿਧੀ

ਸਹਾਇਕ ਪੁਸਤਕਾਂ:

ਸਾਹਿਤ ਦੇ ਰੰਗ (ਸੰਪ. ਡਾ. ਮਹਿਲ ਸਿੰਘ), ਰਵੀ ਸਾਹਿਤ ਪ੍ਰਕਾਸ਼ਨ ਅੰਮ੍ਰਿਤਸਰ।

ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਦਾ ਵਿਆਕਰਣ ਜੋਗਿੰਦਰ ਸਿੰਘ ਪੁਆਰ, ਬਲਦੇਵ ਸਿੰਘ ਚੀਮਾ, ਸੁਖਵਿੰਦਰ ਸਿੰਘ ਸੰਘਾ, ਵੇਦ ਅਗਨੀਹੋਤਰੀ, ਪੰਜਾਬੀ ਭਾਸ਼ਾ ਅਕਾਦਮੀ, ਜਲੰਧਰ, ਐਡੀਸ਼ਨ 2009.

BSNM-23609
Drug Abuse-II (Management and Prevention)

L T P C
2 0 0 0

Course Objective

This course will cover some basic information about the nature and effects of drugs of abuse. The course will also address the social and personal dynamics involved in the phenomena of drug abuse. Finally, the various measures for prevention currently being employed to combat drug abuse will be explored.

I. Prevention of Drug abuse **(6)**

Role of family: Parent child relationship, Family support, Supervision, Shaping values, Active Scrutiny.

II. Prevention of Drug abuse **(6)**

School: Counselling, Teacher as role-model. Parent-teacher-Health Professional Coordination
Random testing on students.

III. Controlling Drug Abuse **(6)**

Media: Restraint on advertisements of drugs, advertisements on bad effects of drugs, Publicity and media, Campaigns against drug abuse, Educational and awareness program.

IV. Legislation **(6)**

NDPs act, Statutory warnings, Policing of Borders, Checking Supply/Smuggling of Drugs, Strict enforcement of laws, Time bound trials.

Course Outcome: Students will be able to:

1. Identify some of the current trends in subject matter.
2. Recognize the impact drugs have on individuals and society as a whole.
3. Describe the action of drugs and the role the drug has had in societies of the past and in the current world.
4. Classify drugs into the drug category to which they belong, as well as identify the schedule the drug is assigned in this country.
5. Evaluate the pros and cons of substance use, society's view of different substances, and the pros and cons of current drug law.

Text Books

1. Ahuja, Ram (2003), Social Problems in India, Rawat Publication, Jaipur.
2. Extent, Pattern and Trend of Drug Use in India, Ministry of Social Justice and Empowerment, Government of India, 2004.
3. Inciardi, J.A. 1981. The Drug Crime Connection. Beverly Hills: Sage Publications.

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4. Kapoor. T. (1985) Drug epidemic among Indian Youth, New Delhi: Mittal Pub. 15
5. Modi, Ishwar and Modi, Shalini (1997) Drugs: Addiction and Prevention, Jaipur: Rawat Publication.
6. National Household Survey of Alcohol and Drug abuse. (2003) New Delhi, Clinical Epidemiological Unit, All India Institute of Medical Sciences, 2004.
7. Sain, Bhim 1991, Drug Addiction Alcoholism, Smoking obscenity New Delhi: Mittal Publications.
8. Singh, Chandra Paul 2000. Alcohol and Dependence among Industrial Workers: Delhi: Shipra.
9. Sussman, S and Ames, S.L. (2008). Drug Abuse: Concepts, Prevention and Cessation, Cambridge University Press.
10. Verma, P.S. 2017, "Punjab's Drug Problem: Contours and Characteristics", Economic and Political Weekly, Vol. LII, No. 3, P.P. 40-43.
11. World Drug Report 2016, United Nations office of Drug and Crime.
12. World Drug Report 2017, United Nations office of Drug and Crime.

BSNM-23610
CHEMISTRY LAB- VI

L	T	P	C
0	0	4	2

Organic Chemistry Laboratory Techniques

(a) Column Chromatography Separation of o & p-nitrophenol Separation of Leaf pigments from Spinnach leaves Separation of o & p-nitro aniline Separation of dyes

(b) **Synthesis of Organic Compounds** Preparation of p-nitroacetanilide Preparation of p-romoacetanilide

Green Chemistry Experiment: Preparation of benzilic acid from Benzyl-using green approach.
Preparation of Methyl Orange, Methyl Red Preparation of benzilic acid from benzyl-using green approach.

Text Books

1. Experimental Organic Chemistry, Vol. I & II, P.R. Singh, D.S. Gupta and K.S. Bajpai, Tata McGraw Hill.
2. Laboratory Manual in Organic Chemistry, R.K. Bansal, Wiley Eastern.
3. Vogel's Textbook of Practical Organic Chemistry, B.S. Furniss, A.J. Hannaford, V. Rogers, P.W.G. Smith and A.R. Tatchell, ELBS.
4. Practical Organic Chemistry by F.G. Mann and B.C. Saunders

BSNM-23611
PHYSICS LAB- VI

L	T	P	C
0	0	4	2

Students are required to perform at least 06 experiments from the following:

1. Characteristics of pn junction diode
2. Characteristics of Zener diode.
3. To determine the resistivity of semiconductors.
4. Measurement of susceptibility of paramagnetic solution (Quinck`s Tube Method)
5. To measure the Magnetic susceptibility of Solids.
6. To determine the Coupling Coefficient of a Piezoelectric crystal.
7. To measure the Dielectric Constant of a dielectric Materials with frequency.
8. To determine the complex dielectric constant and plasma frequency of metal using Surface Plasmon resonance (SPR).
9. To determine the refractive index of a dielectric layer using SPR.
10. To study the PE Hysteresis loop of a Ferroelectric Crystal.
11. To study the BH curve of iron using a Solenoid and determine the energy loss.
12. To measure the resistivity of a semiconductor (Ge) crystal with temperature by four-probe method (room temperature to 150°C and to determine its band gap.
13. To determine the Hall coefficient of a semiconductor sample.

Text Books

1. Advanced Practical Physics for students, B. L. Flint and H.T. Worsnop, 1971, Asia Publishing House.
2. Advanced level Physics Practicals, Michael Nelson and Jon M. Ogborn, 4th Edition, reprinted 1985, Heinemann Educational Publishers
3. Engineering Practical Physics, S.Panigrahi & B. Mallick, 2015, Cengage Learning India Pvt. Ltd.
4. Practical Physics, G.L. Squires, 2015, 4th Edition, Cambridge University Press.
5. A Text Book of Practical Physics, I.Prakash & Ramakrishna, 11th Edn, 201, Kitab Mahal.
6. B Sc. Practical Physics, C. L. Arora, S. Chand & Co.
7. A Text Book of Practical Physics, Indu Prakash and Ramakrishna, 11th Edition, 2011, Kitab Mahal, New Delhi.